



PARADIM

AN NSF MATERIALS INNOVATION PLATFORM

LECTURE #5— SUBOXIDE MBE

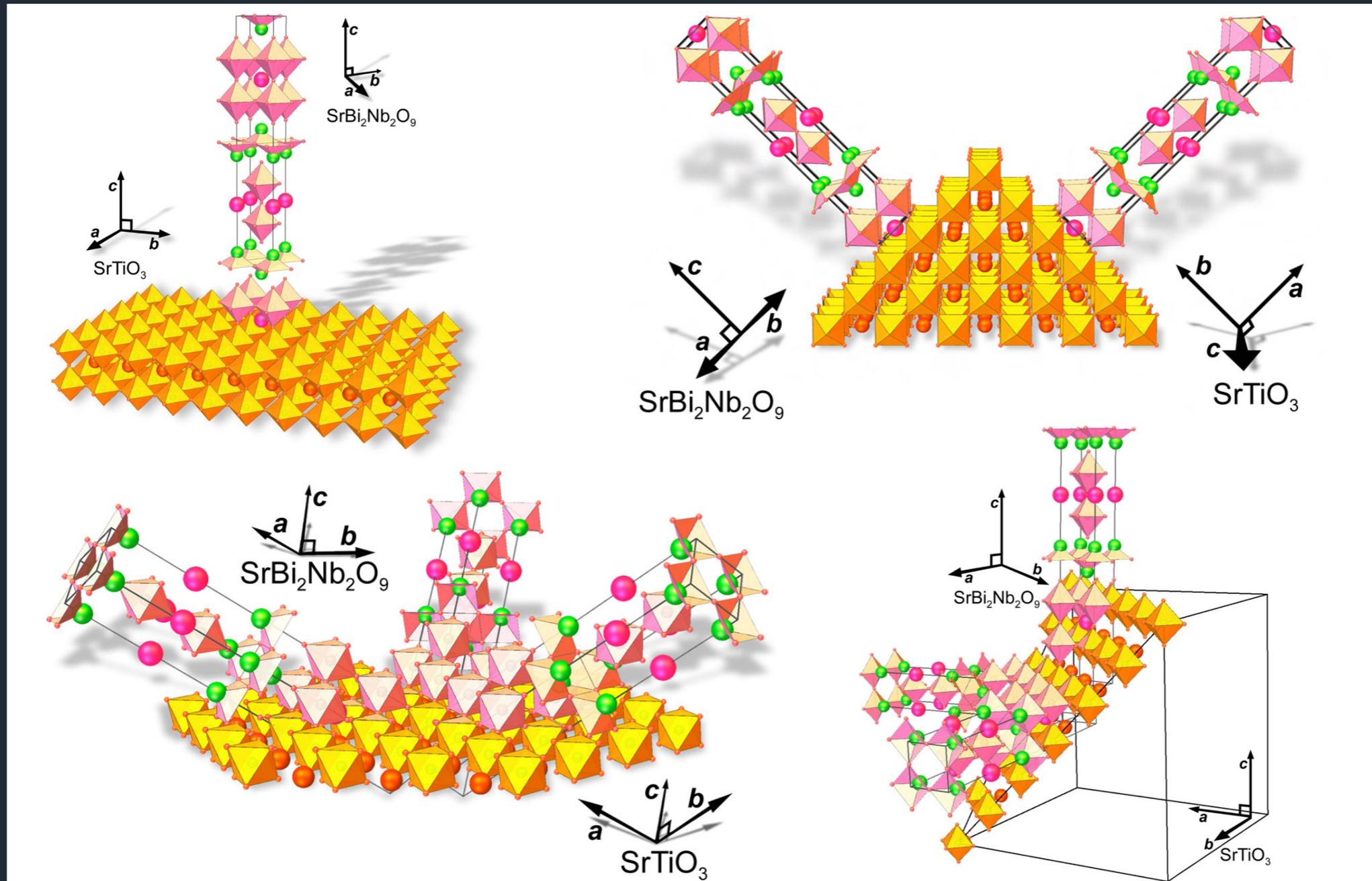
Darrell G. Schlom

*Department of Materials Science and Engineering
Cornell University*

Kavli Institute at Cornell for Nanoscale Science

Leibniz-Institut für Kristallzüchtung

Epitaxial Growth



D.G. Schlom, L.Q. Chen,
X.Q. Pan, A. Schmehl, and
M.A. Zurbuchen,
*Journal of the American
Ceramic Society* 91 (2008)
2429-2454.

RHEED of $\text{YBa}_2\text{Cu}_3\text{O}_7$

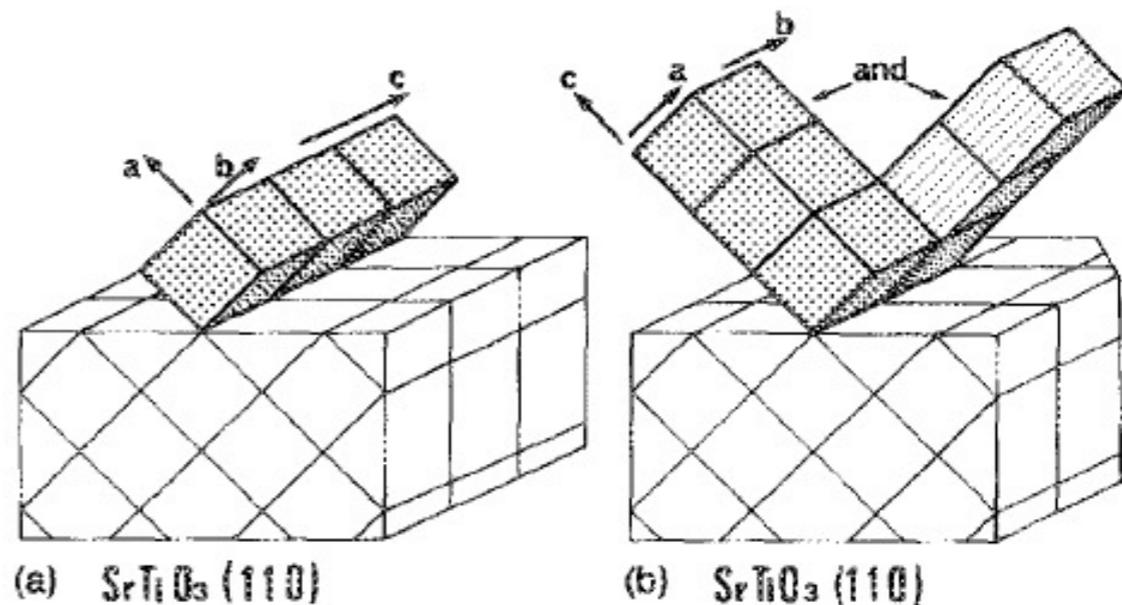
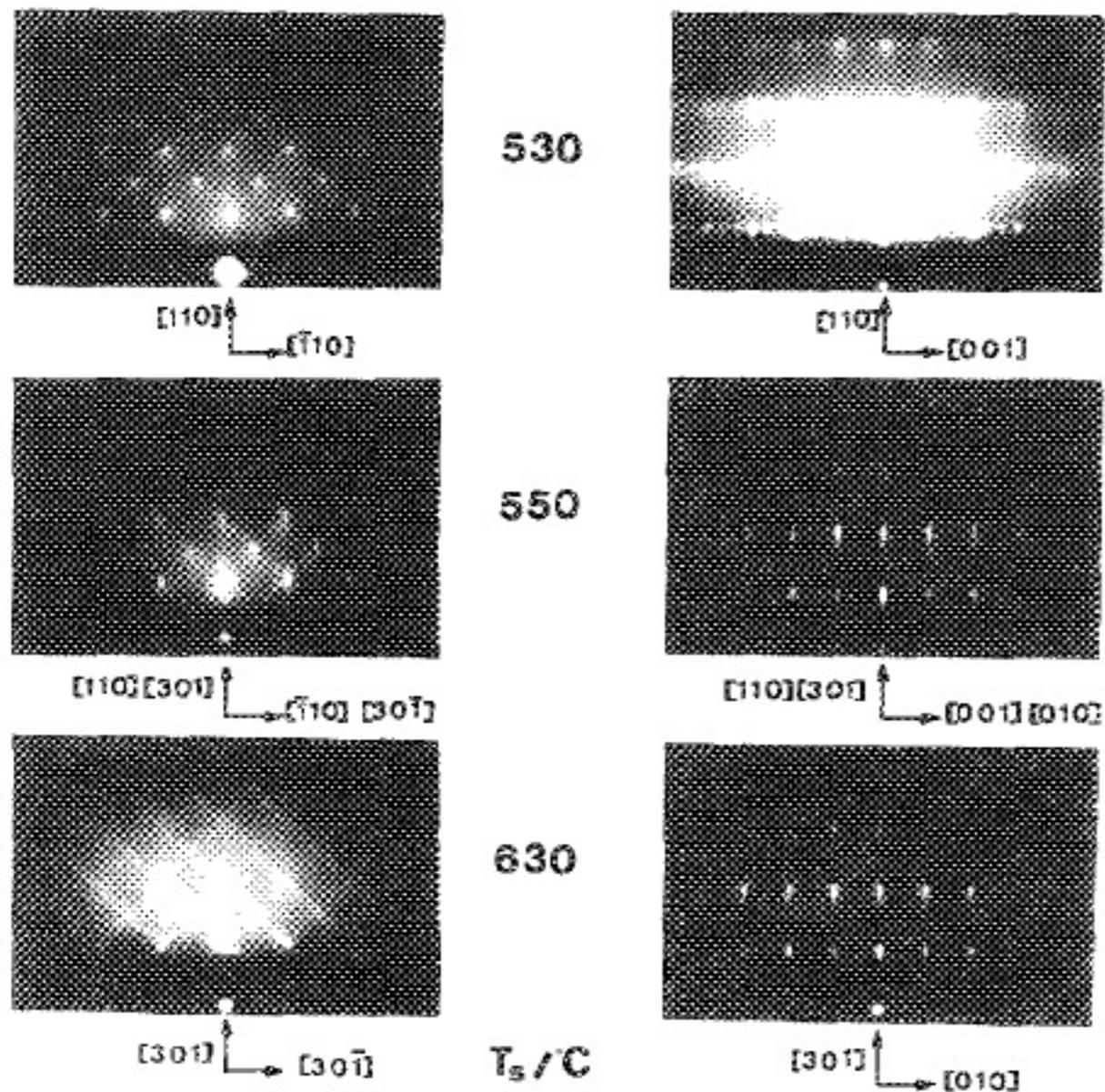
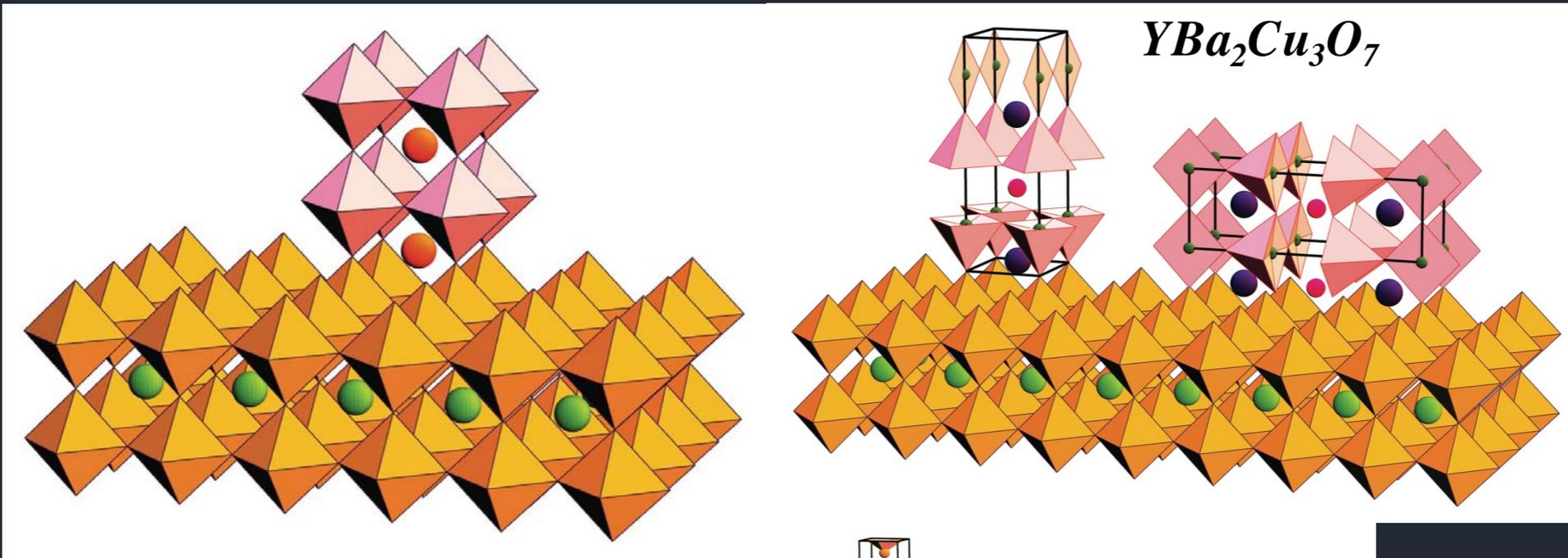


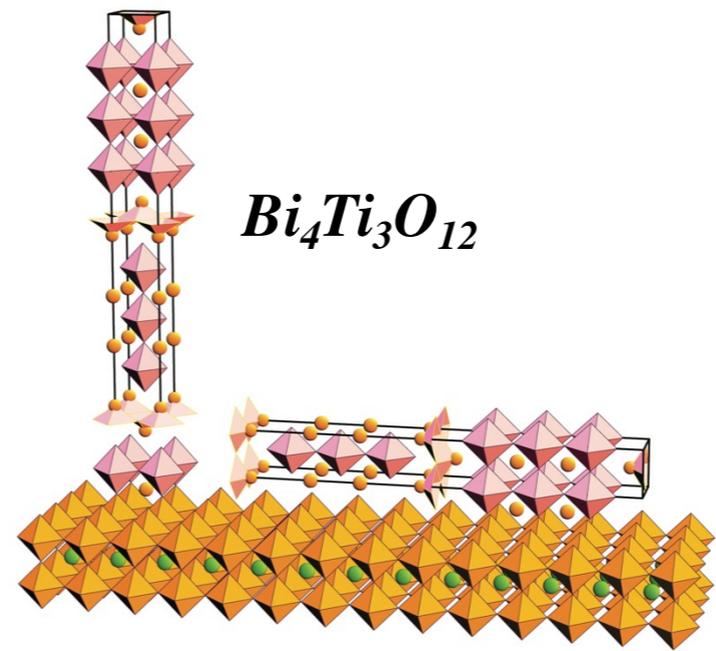
FIG. 1. Reflection high-energy electron diffraction (RHEED) patterns for the YBCO films grown at various temperatures.

T. Terashima, Y. Bando, K. Iijima, K. Yamamoto, K. Hirata, "Epitaxial Growth of $\text{YBa}_2\text{Cu}_3\text{O}_{7-x}$ Thin Films on (110) SrTiO_3 Single Crystals by Activated Reactive Evaporation," *Appl Phys Lett.* 53 (1988) 2232-2234.

Epitaxial Growth



ABO_3 / $A'B'O_3$
“Cube-on-Cube”

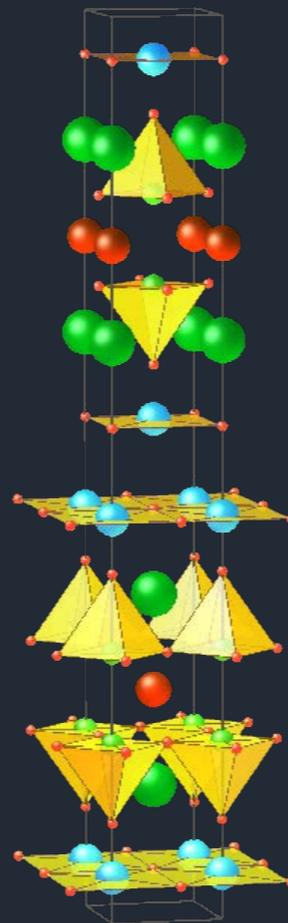
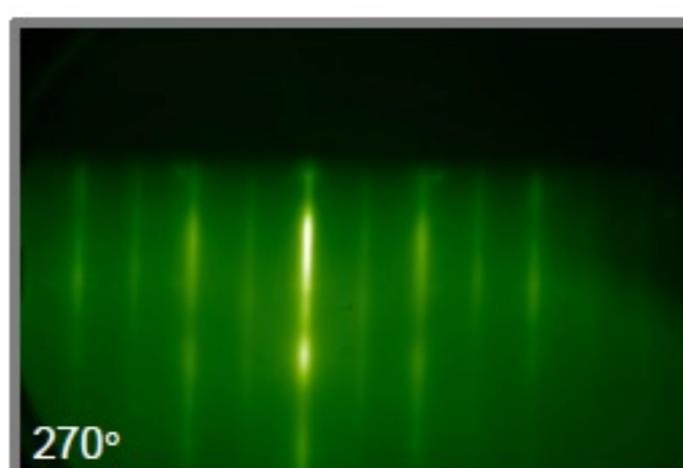
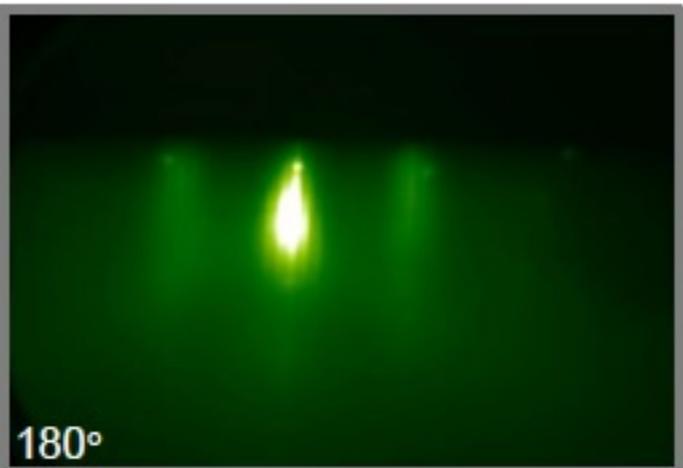
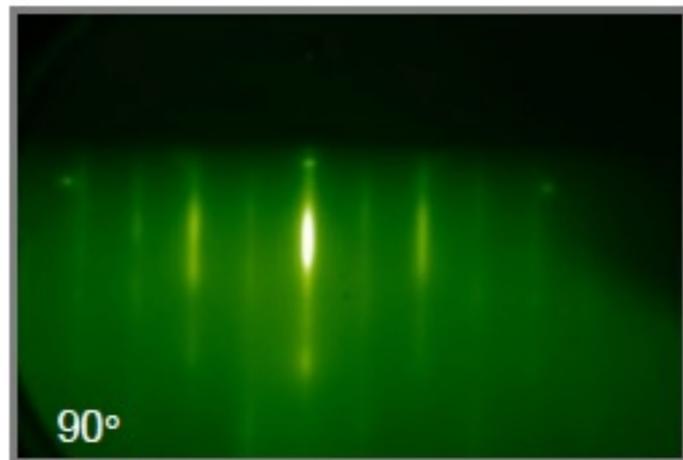


RHEED of $\text{Bi}_2\text{Sr}_2\text{CaCu}_2\text{O}_{8+\delta}$ (Bi-2212)



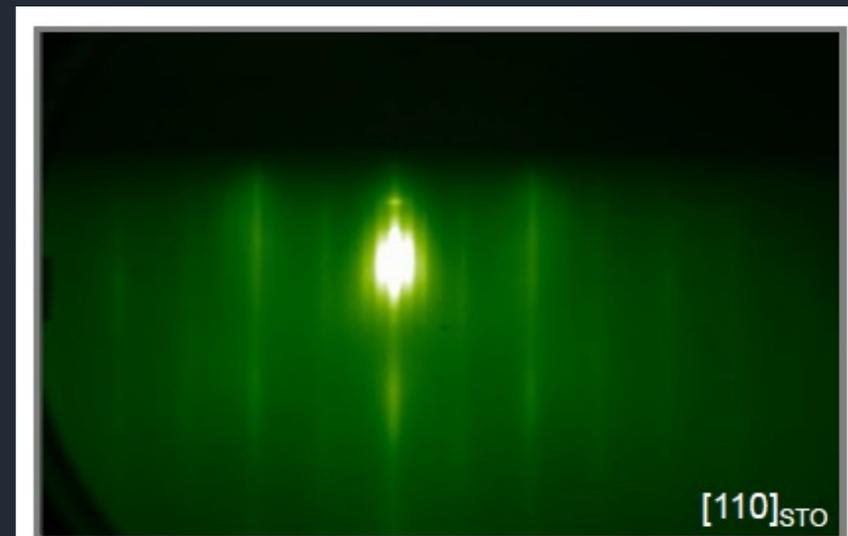
untwinned film

2° miscut; $\langle 110 \rangle$ azimuth



twinned film

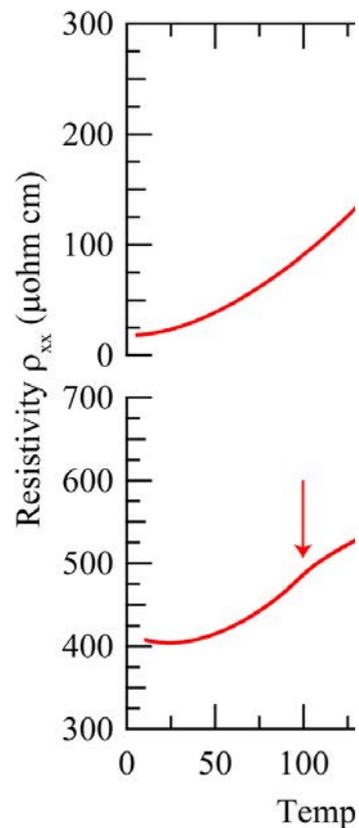
(90° in-plane rotation twins)



Benc

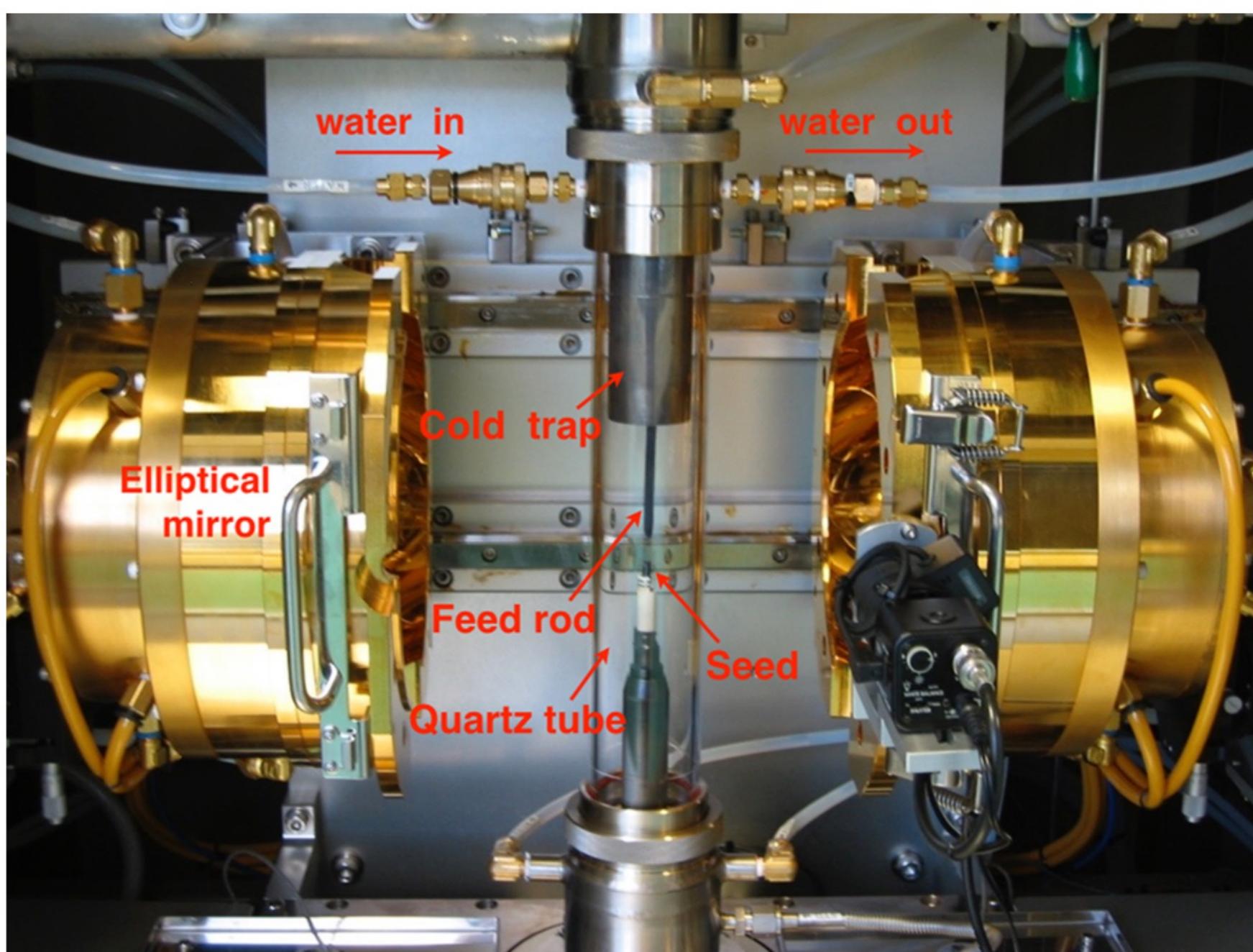
Best PL

$\rho_{300\text{ K}} / \rho_{10\text{ K}}$



~20 nm SrRuO₃

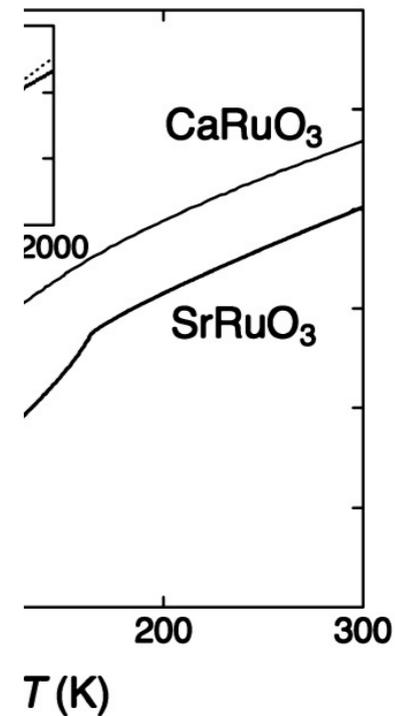
D. Kan, R. Aso, H. Kura
J. Appl. Phys. 113



S

Single Crystal

$\rho_{300\text{ K}} / \rho_{10\text{ K}} = 192$



Single crystal

ambach, J.S. Brooks,
ji, and Y. Maeno,
15 (2015) 5573-5577.

Figure 1. Photograph of an infrared image furnace. In this study, we employed a cold trap in which water is circulating during the growth.

Outline of MBE Lectures



- What is MBE and what is it good for?
Lecture #1 – Greatest hits of MBE
- How to grow your favorite oxide by MBE?
Lectures #2-4 – Nuts and bolts of oxide MBE
- Detailed Examples of Oxide MBE
Lectures #5,6 – **Suboxide MBE**
High Purity Synthesis of Binary Oxides
- How can I gain access to an oxide MBE if I don't have one?
Use PARADIM's oxide MBE (+ ARPES + ...)

Why Ga₂O₃?

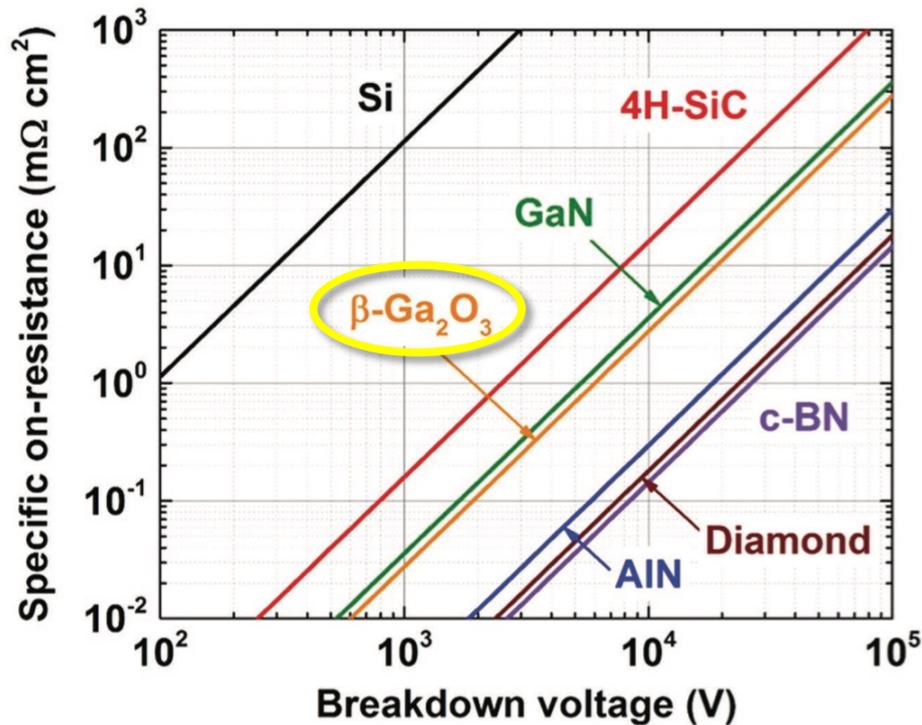


Figure 1. Contours of constant Baliga figure-of-merit (BFOM) for various conventional, WBG and UWBG semiconductors, drawn on a log-log specific on-resistance versus breakdown voltage plot. This is the figure-of-merit of interest for low-frequency unipolar vertical power switches; the lower right region represents higher BFOM, hence higher performance.

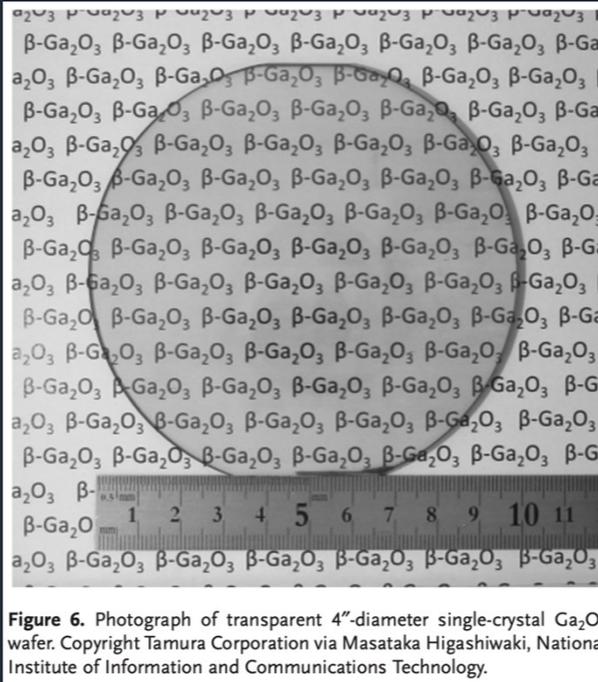


Figure 6. Photograph of transparent 4"-diameter single-crystal Ga₂O₃ wafer. Copyright Tamura Corporation via Masataka Higashiwaki, National Institute of Information and Communications Technology.

4" β-Ga₂O₃ Single-Crystal Substrate

β-Ga₂O₃ has High Bandgap

$$E_g = 4.7 \text{ eV}$$

High Breakdown Field

$$E_{\text{max}} \approx 5 \text{ MV/cm}$$

High Baliga Figure of Merit

Dopable *n*-Type with good mobility

$$E_D \approx 0.02 \text{ eV (Si)}$$

$$\mu \approx 200 \text{ cm}^2/(\text{V}\cdot\text{s})$$

Thermal Conductivity

$$k = 11\text{-}29 \text{ W/(m K)}$$

J.Y. Tsao, S. Chowdhury, M.A. Hollis, D. Jena, N.M. Johnson, K.A. Jones, R.J. Kaplar, S. Rajan, C.G.V. de Walle, E. Bellotti, C.L. Chua, R. Collazo, M.E. Coltrin, J.A. Cooper, K.R. Evans, S. Graham, T.A. Grotjohn, E.R. Heller, M. Higashiwaki, M.S. Islam, P.W. Juodawlkis, M.A. Khan, A.D. Koehler, J.H. Leach, U.K. Mishra, R.J. Nemanich, R.C.N. Pilawa-Podgurski, J.B. Shealy, Z. Sitar, M.J. Tadjer, A.F. Witulski, M. Wraback, and J.A. Simmons, "Ultrawide-Bandgap Semiconductors: Research Opportunities and Challenges," *Advanced Electronic Materials* 4 (2018) 1600501.

Conventional (Ga) MBE of β -Ga₂O₃



Low Growth Rate

- Growth Rate: 0.2 $\mu\text{m/hr}$ (maximum reported is 0.7 $\mu\text{m/hr}$)
- Peak Mobility at this Growth Rate: 120 $\text{cm}^2/(\text{V}\cdot\text{s})$ at room temperature

E. Ahmadi, O.S. Koksaldi, S.W. Kaun, Y. Oshima, D.B. Short, U.K. Mishra, and J.S. Speck,
“Ge Doping of β -Ga₂O₃ Films Grown by Plasma-Assisted Molecular Beam Epitaxy,” *Applied Physics Express* 10 (2017) 041102.

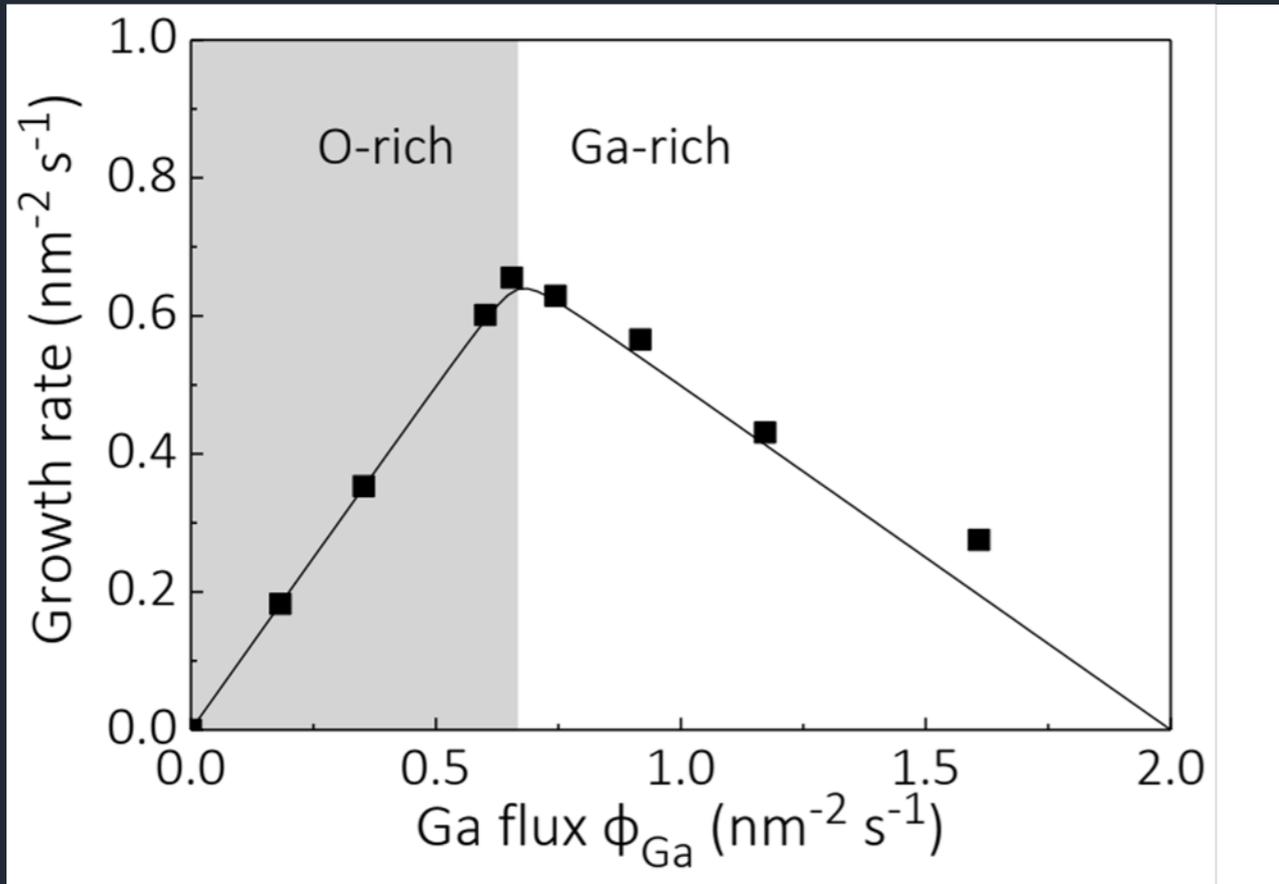
MOCVD is considerably better

- Growth Rate: ~0.5 $\mu\text{m/hr}$ (up to 10 $\mu\text{m/hr}$ reported)
- Peak Mobility at this Growth Rate: 194 $\text{cm}^2/(\text{V}\cdot\text{s})$ at room temperature

Z. Feng, A.F.M.A.U. Bhuiyan, Z. Xia, W. Moore, Z. Chen, J.F. McGlone, D.R. Daughton, A.R. Arehart, S.A. Ringel, S. Rajan, and H. Zhao,
“Probing Charge Transport and Background Doping in Metal-Organic Chemical Vapor Deposition-Grown (010) β -Ga₂O₃,”
Physica Status Solidi RRL 14 (2020) 2000145.

Can MBE be improved for the growth of β -Ga₂O₃?

Conventional (Ga) MBE of β -Ga₂O₃

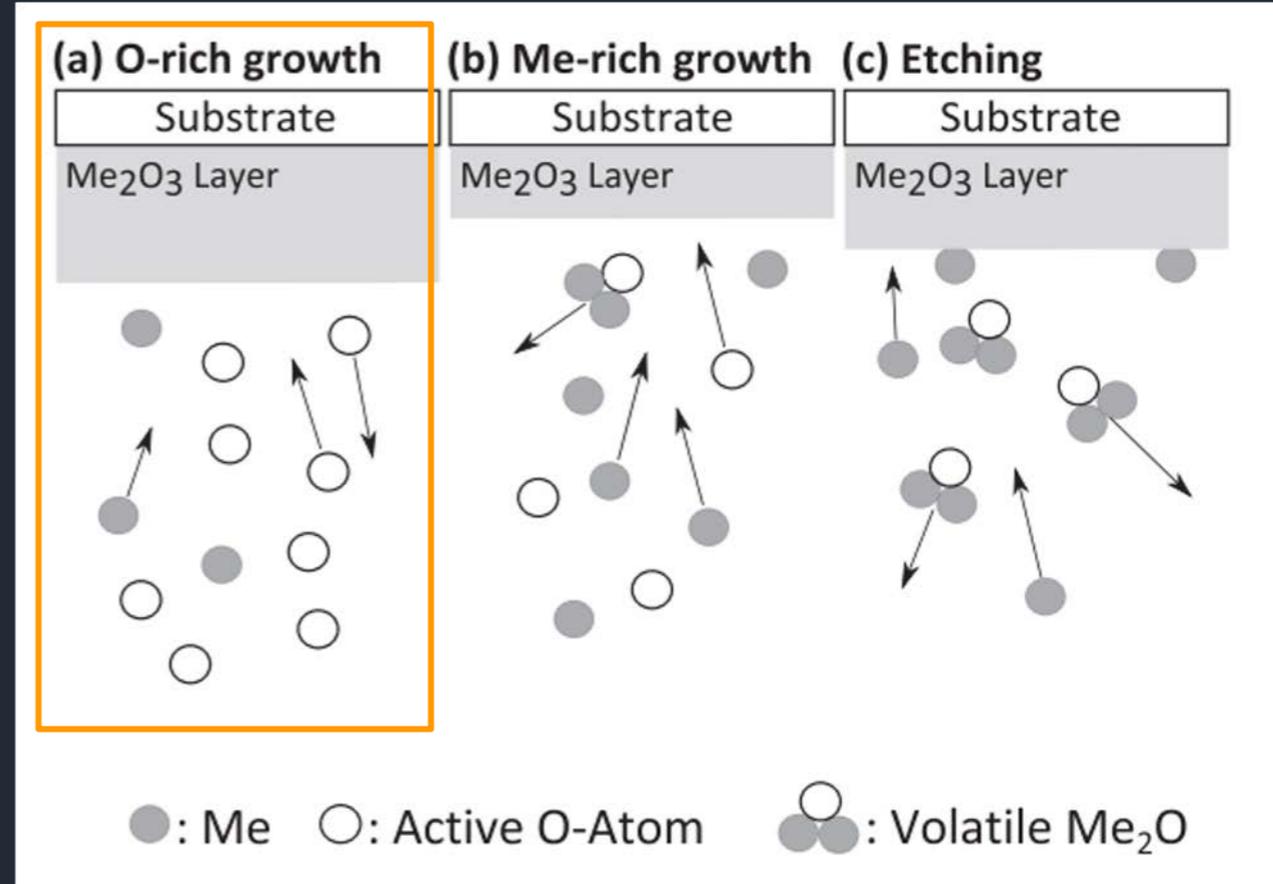
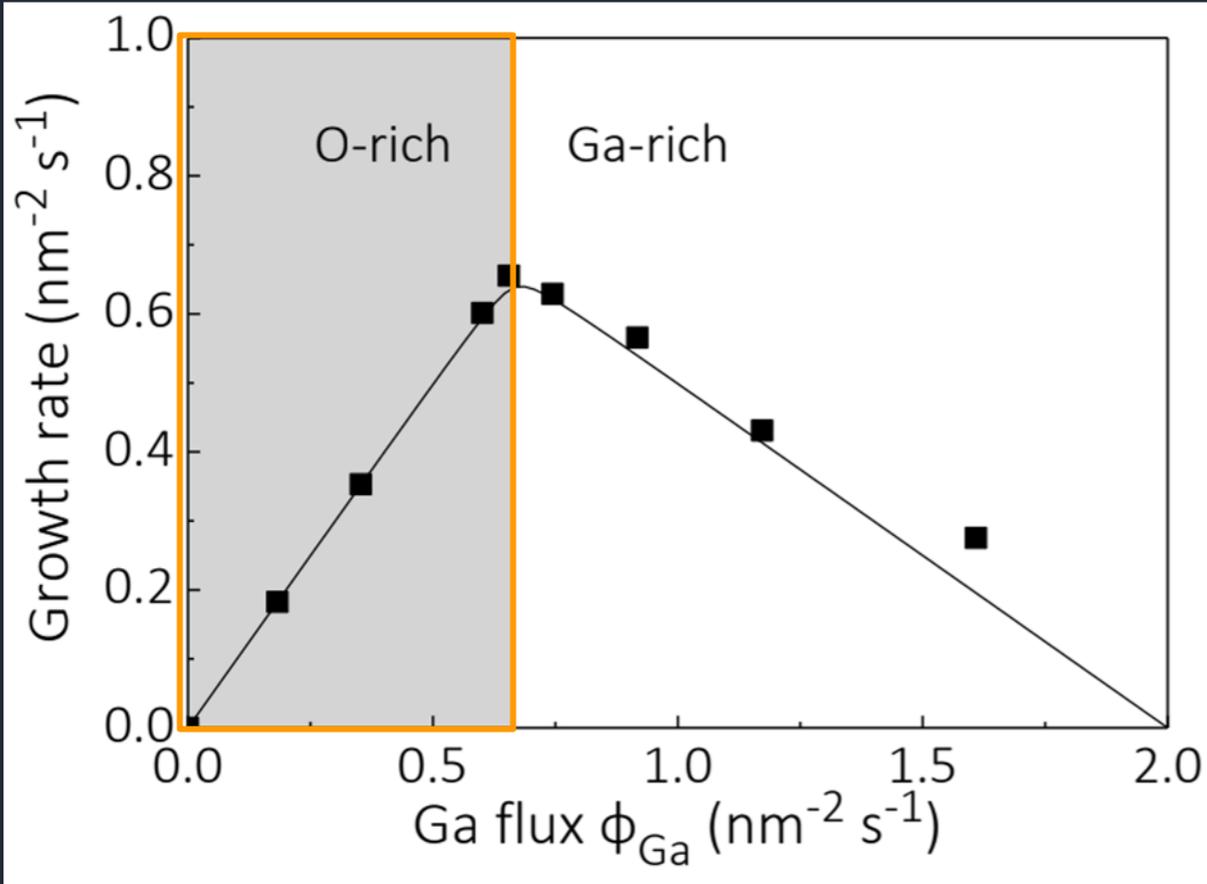


P. Vogt and O. Bierwagen,

“The Competing Oxide and Sub-Oxide Formation in Metal-Oxide Molecular Beam Epitaxy,”

Applied Physics Letters 106 (2015) 081910.

Conventional (Ga) MBE of $\beta\text{-Ga}_2\text{O}_3$



2-step reaction mechanism explains:



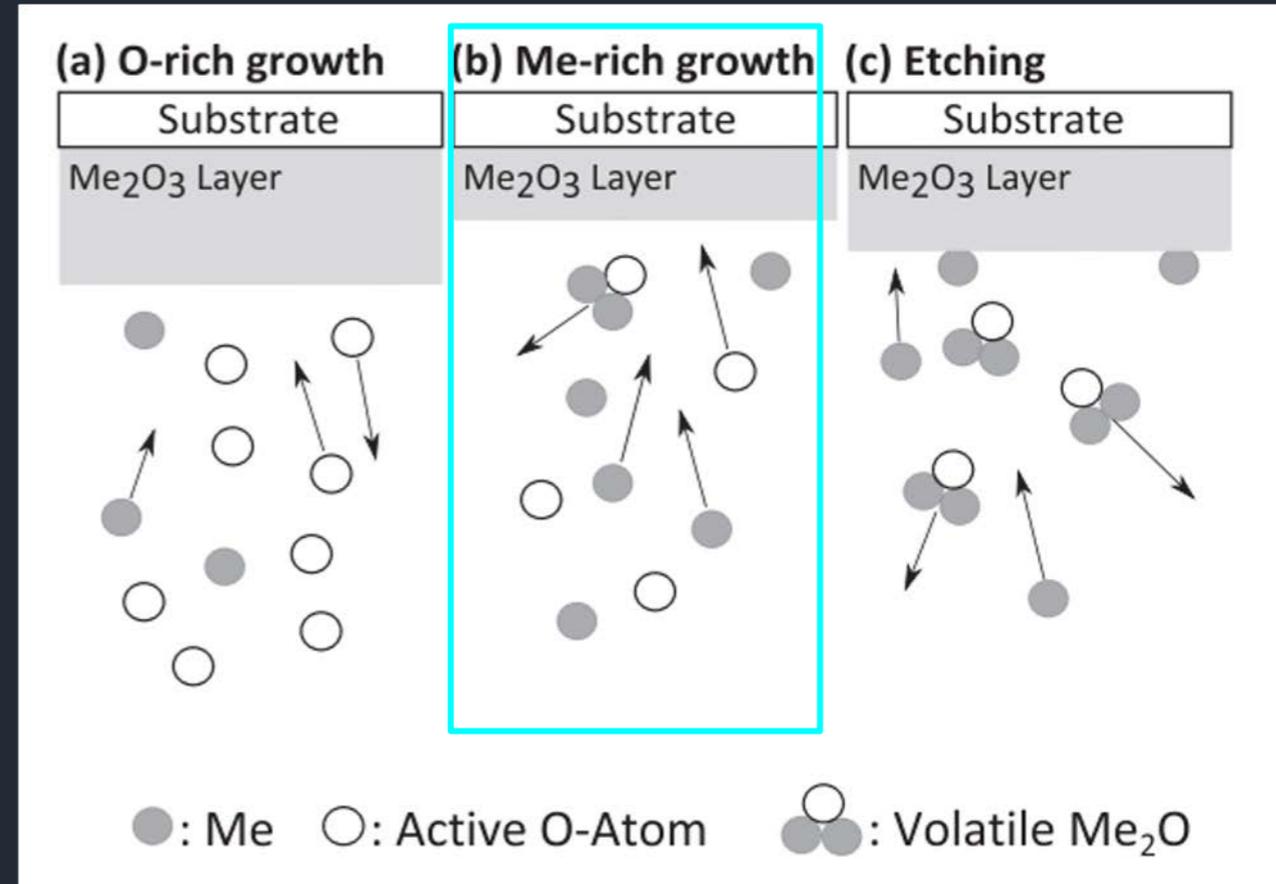
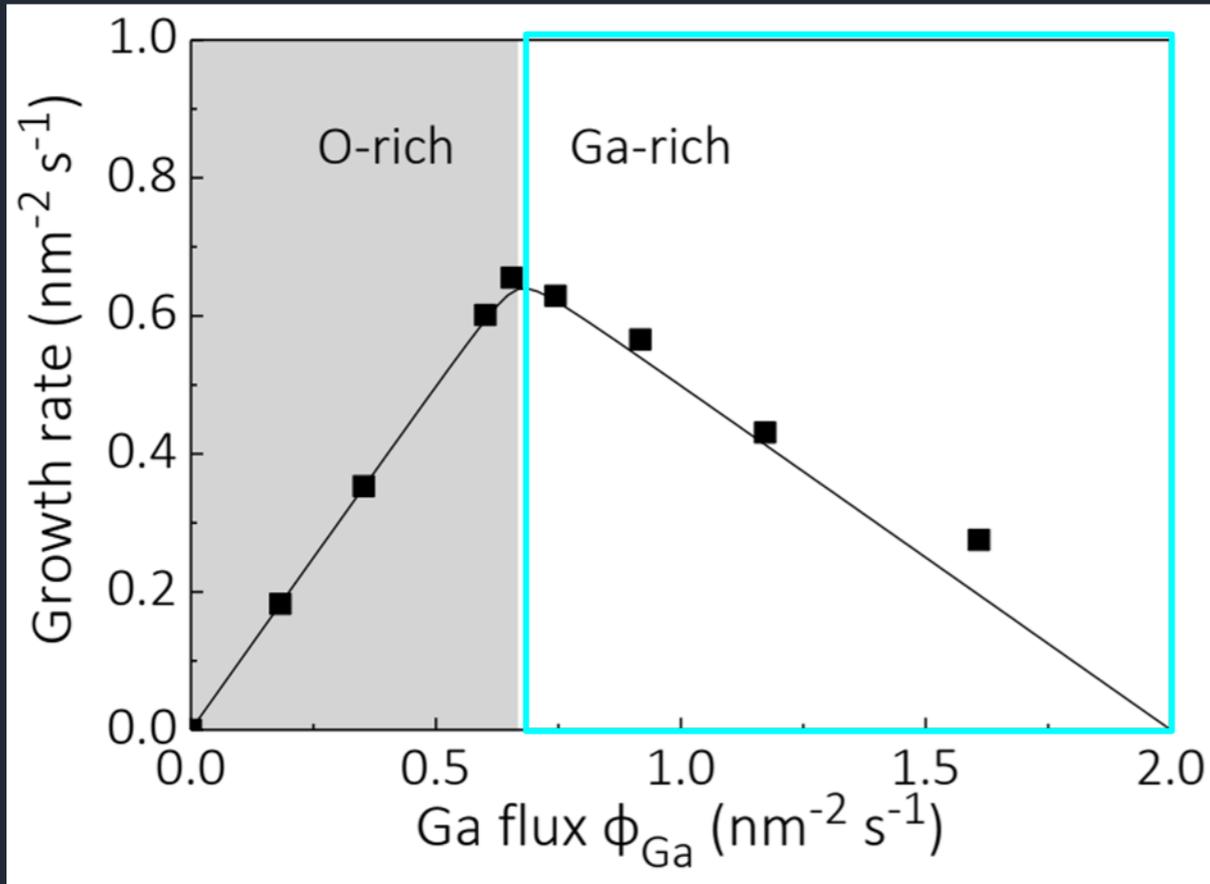
P. Vogt and O. Bierwagen,

“The Competing Oxide and Sub-Oxide Formation in Metal-Oxide Molecular Beam Epitaxy,”
Applied Physics Letters 106 (2015) 081910.

P. Vogt and O. Bierwagen,

“Quantitative Subcompound-Mediated Reaction Model for the Molecular Beam Epitaxy of III-VI and IV-VI Thin Films: Applied to Ga_2O_3 , In_2O_3 , and SnO_2 ,”
Physical Review Materials 2 (2018) 120401.

Conventional (Ga) MBE of $\beta\text{-Ga}_2\text{O}_3$



2-step reaction mechanism explains:



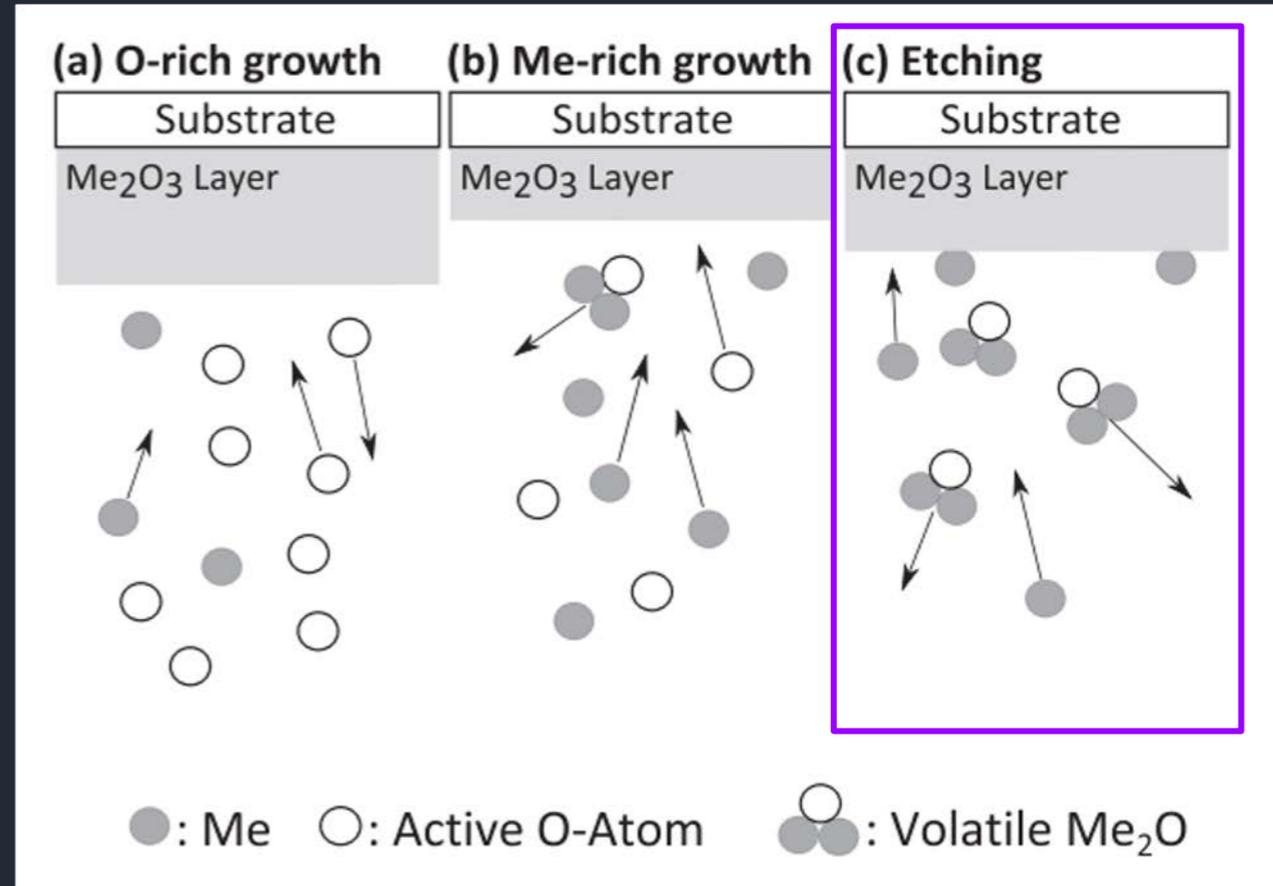
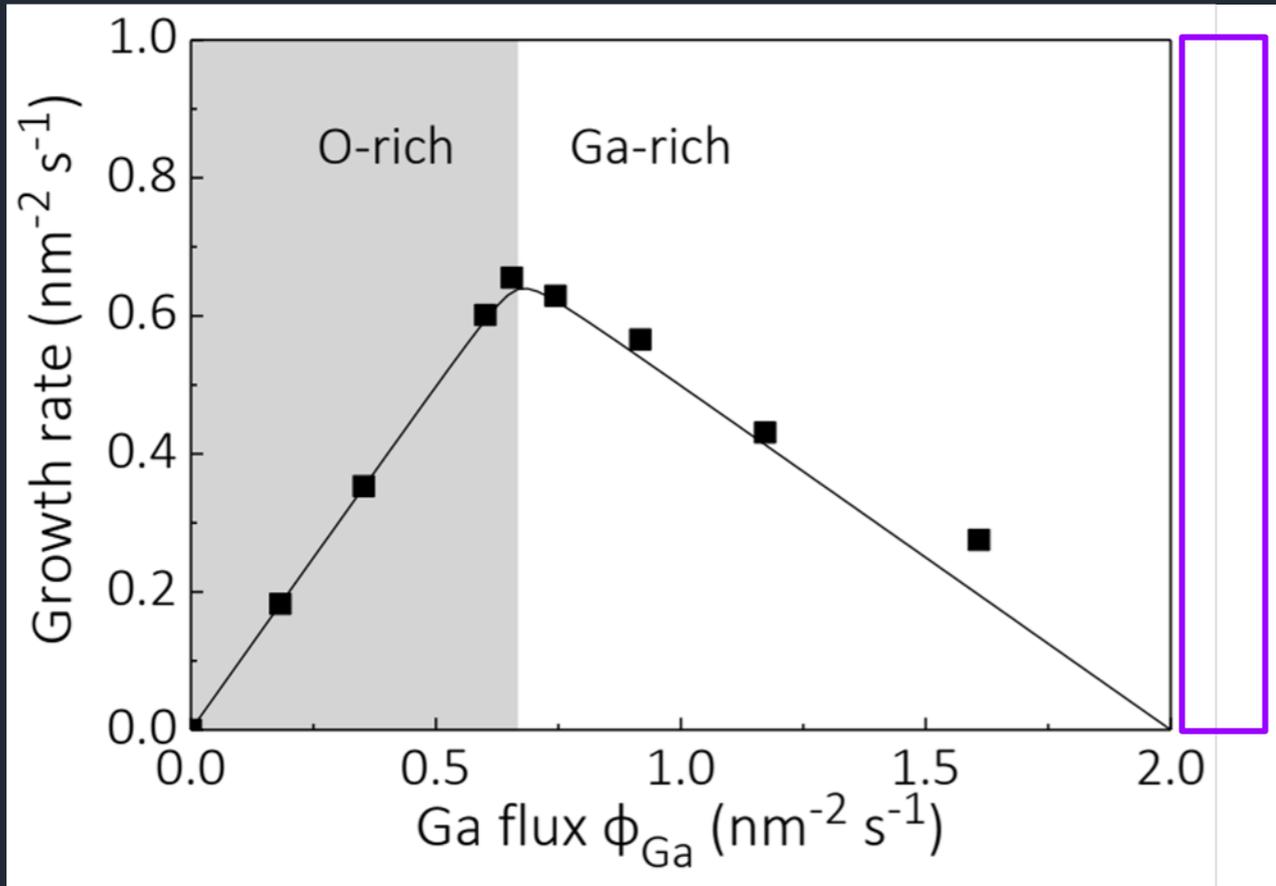
P. Vogt and O. Bierwagen,

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Applied Physics Letters 106 (2015) 081910.

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“Quantitative Subcompound-Mediated Reaction Model for the Molecular Beam Epitaxy of III-VI and IV-VI Thin Films: Applied to Ga₂O₃, In₂O₃, and SnO₂,”
Physical Review Materials 2 (2018) 120401.

Conventional (Ga) MBE of β -Ga₂O₃



2-step reaction mechanism explains:



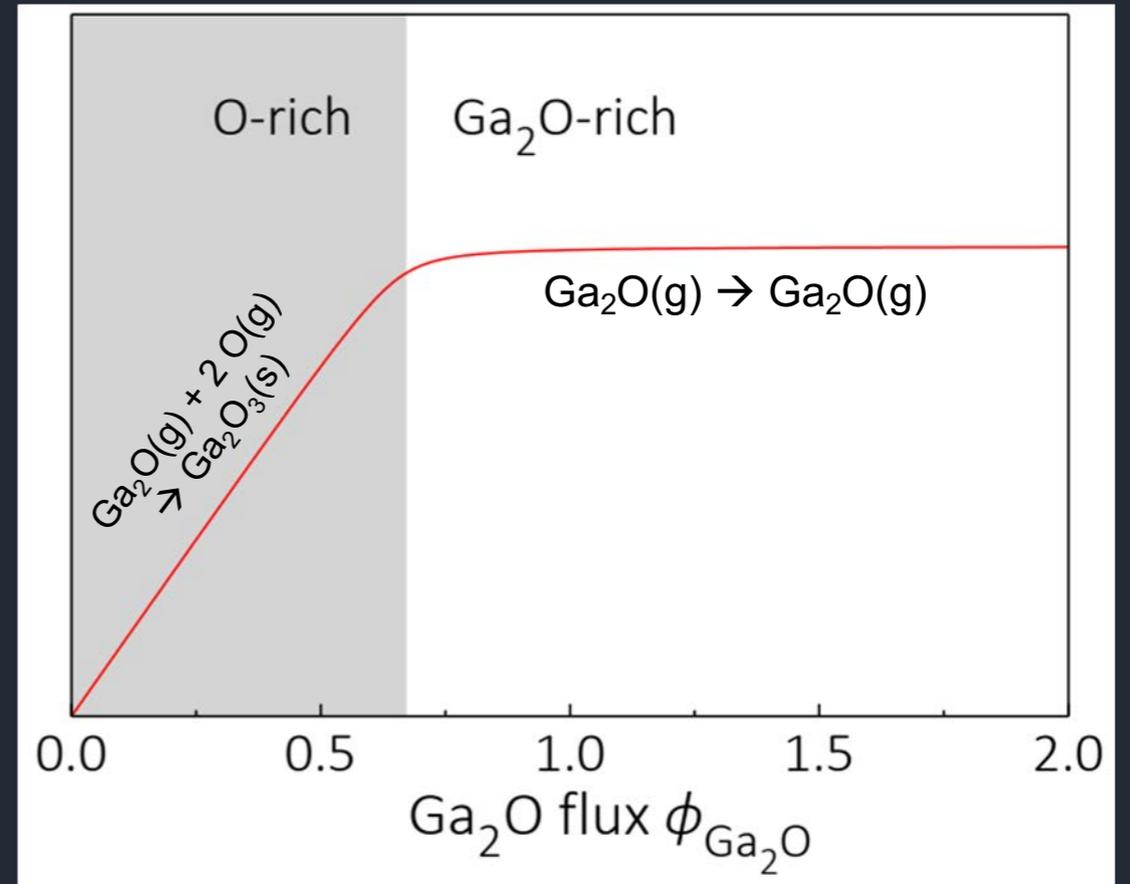
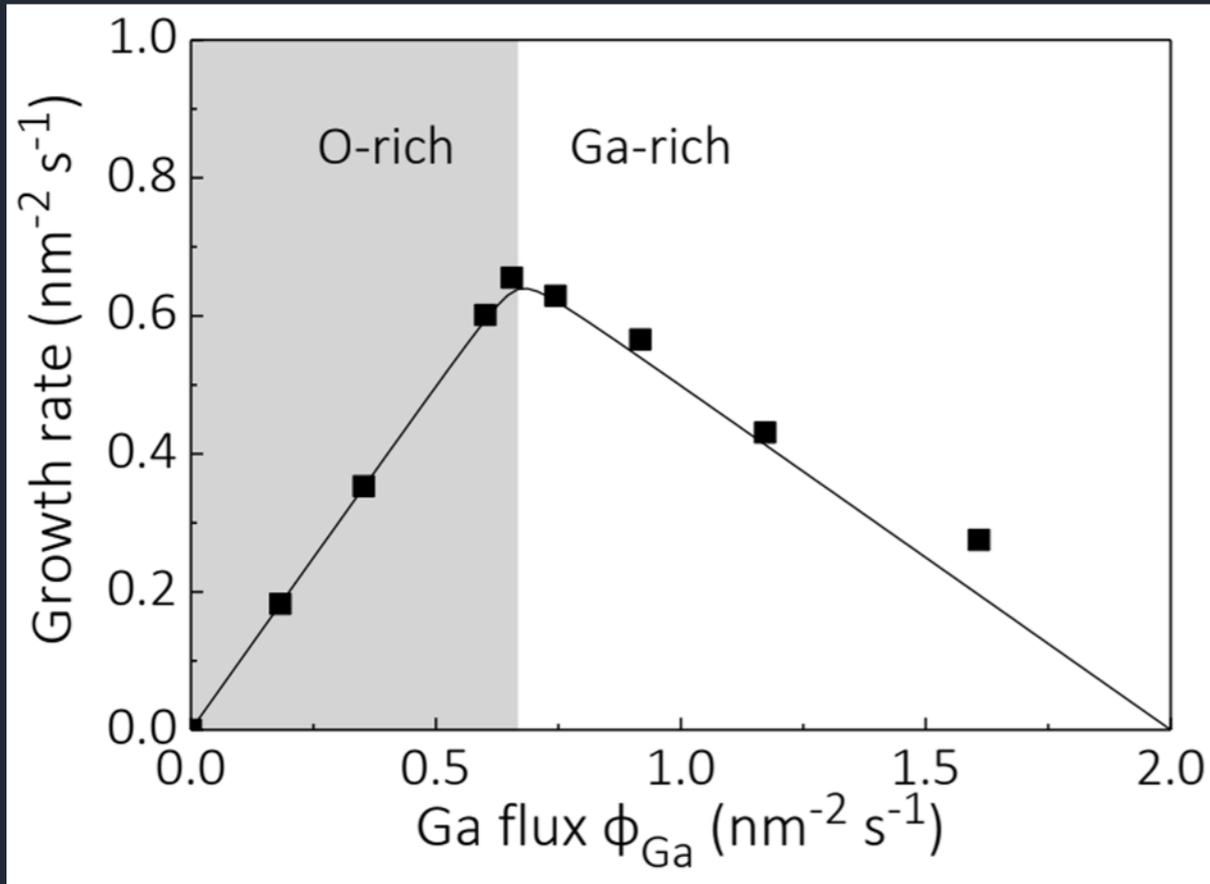
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Applied Physics Letters 106 (2015) 081910.

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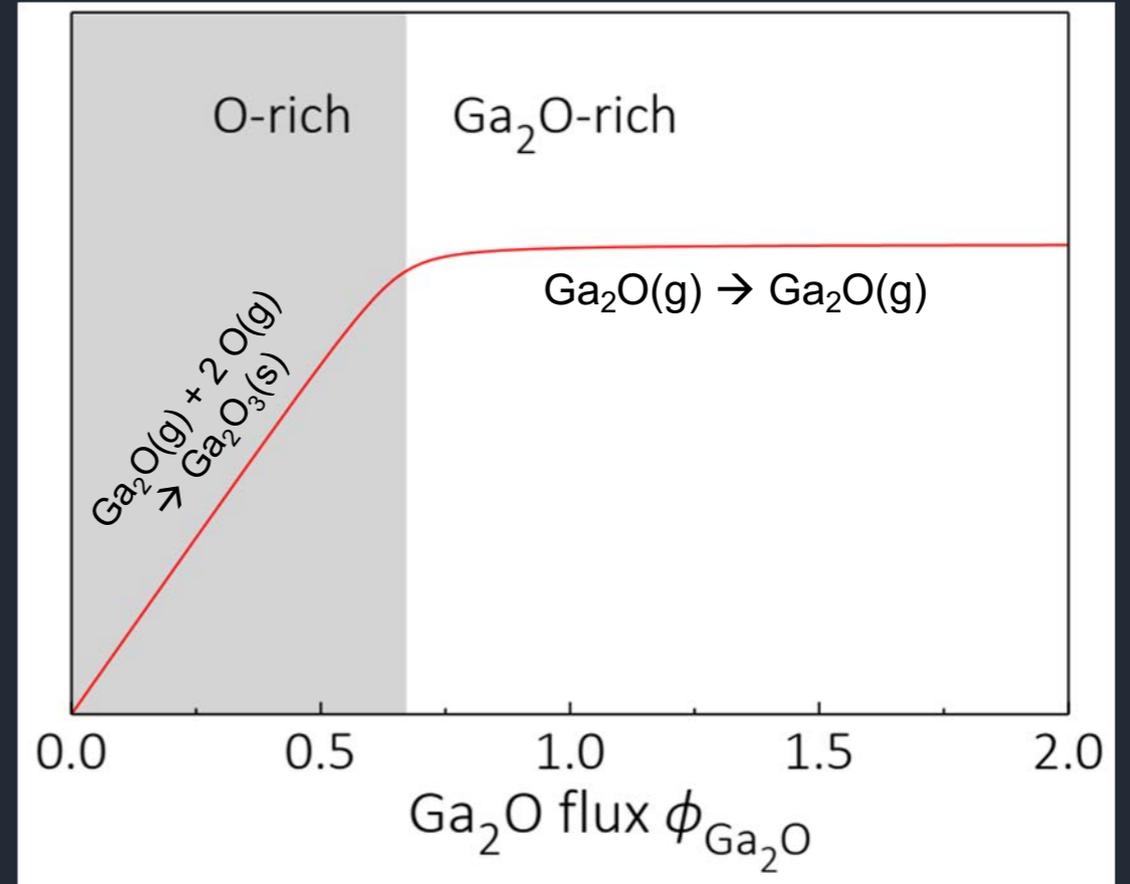
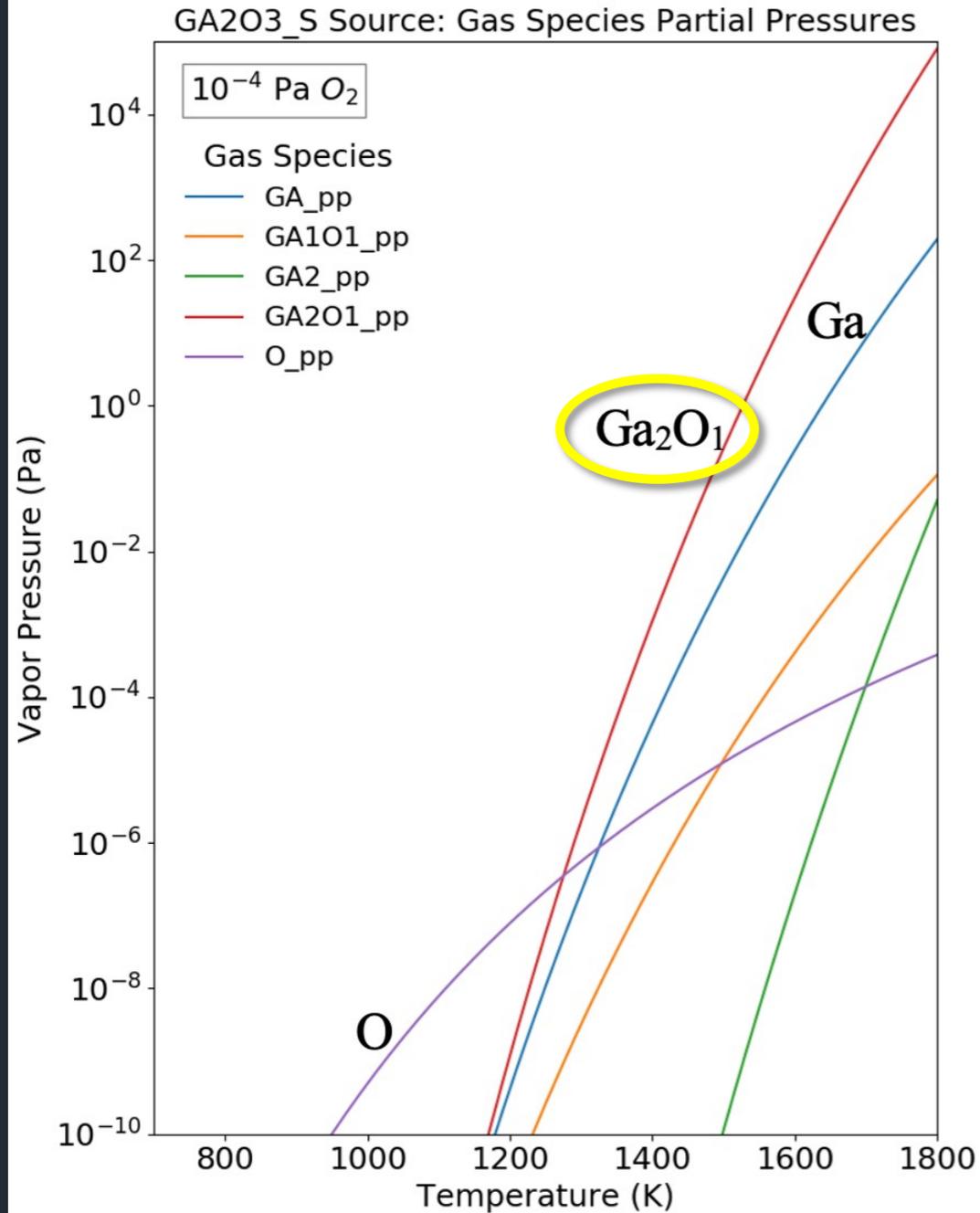
Suboxide (Ga_2O) MBE of $\beta\text{-Ga}_2\text{O}_3$



2-step reaction mechanism explains:

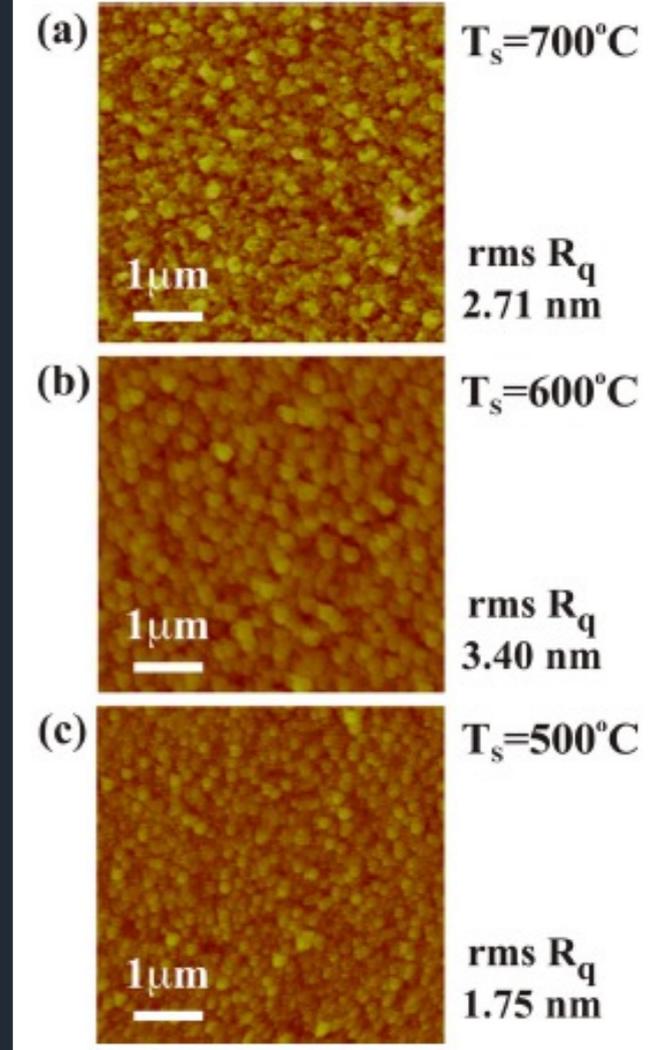
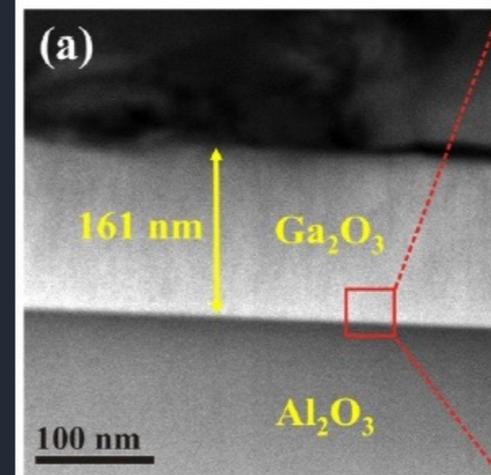
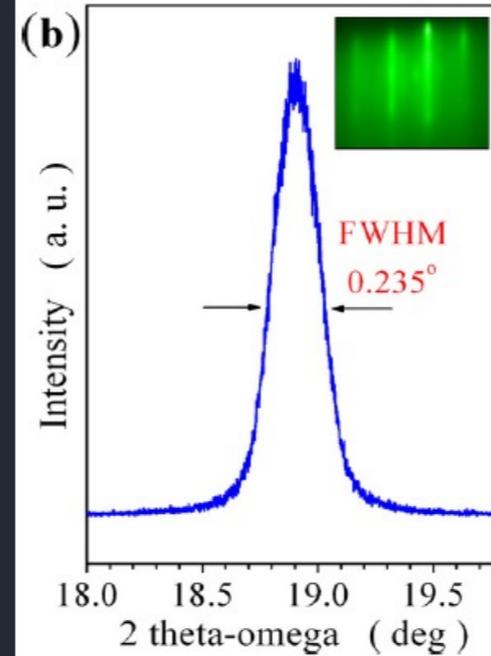
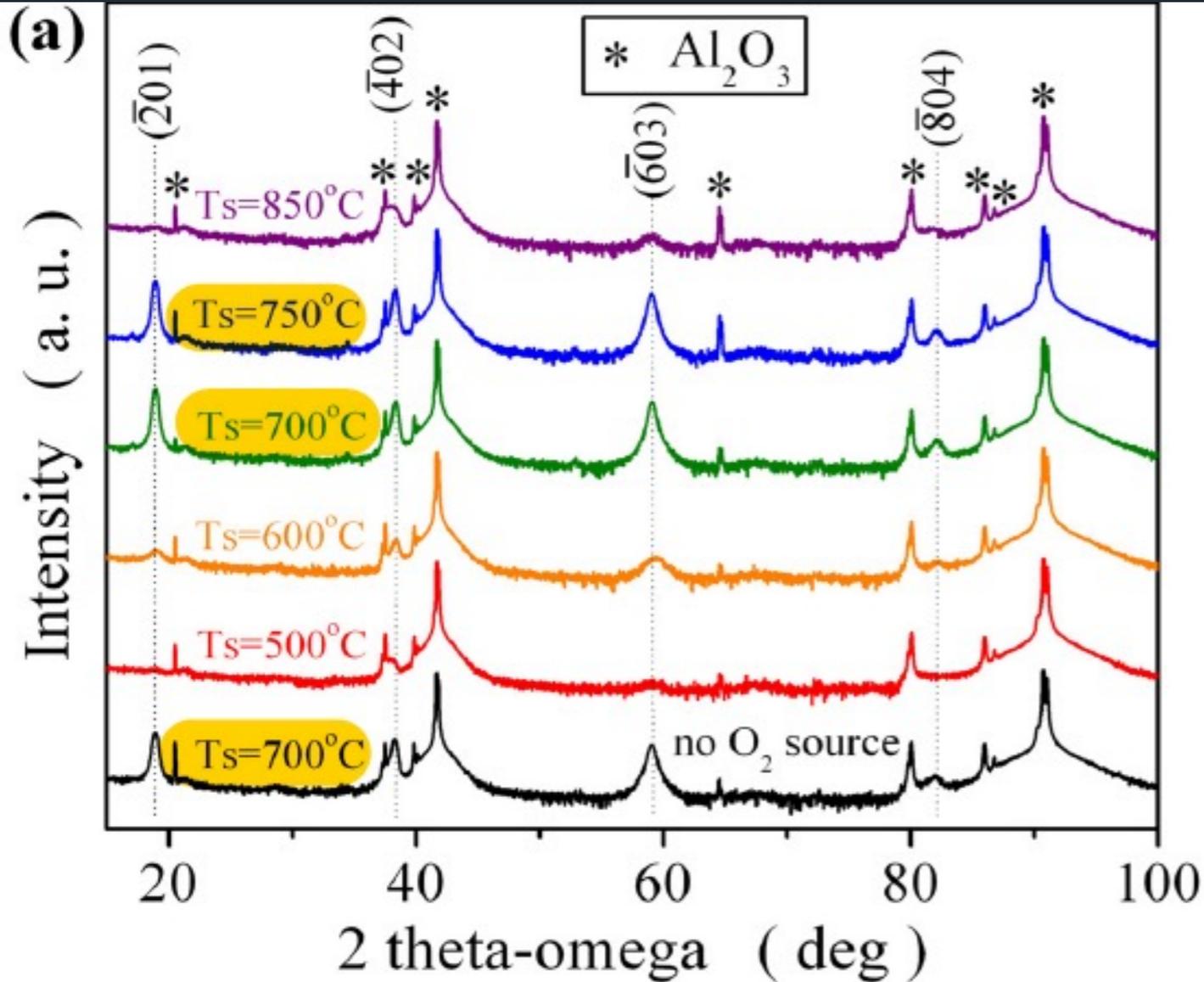


Use $\text{Ga}_2\text{O}_3(\text{s})$ rather than $\text{Ga}(\ell)$?



K.M. Adkison, S-L. Shang, B.J. Bocklund, D. Klimm, D.G. Schlom, and Z.K. Liu, "Suitability of Binary Oxides for Molecular-Beam Epitaxy Source Materials: A Comprehensive Thermodynamic Analysis," *APL Materials* 8 (2020) 081110.

Use $\text{Ga}_2\text{O}_3(\text{s})$ rather than $\text{Ga}(\ell)$?



Use $\text{Ga}_2\text{O}_3(\text{s})$ rather than $\text{Ga}(\ell)$?

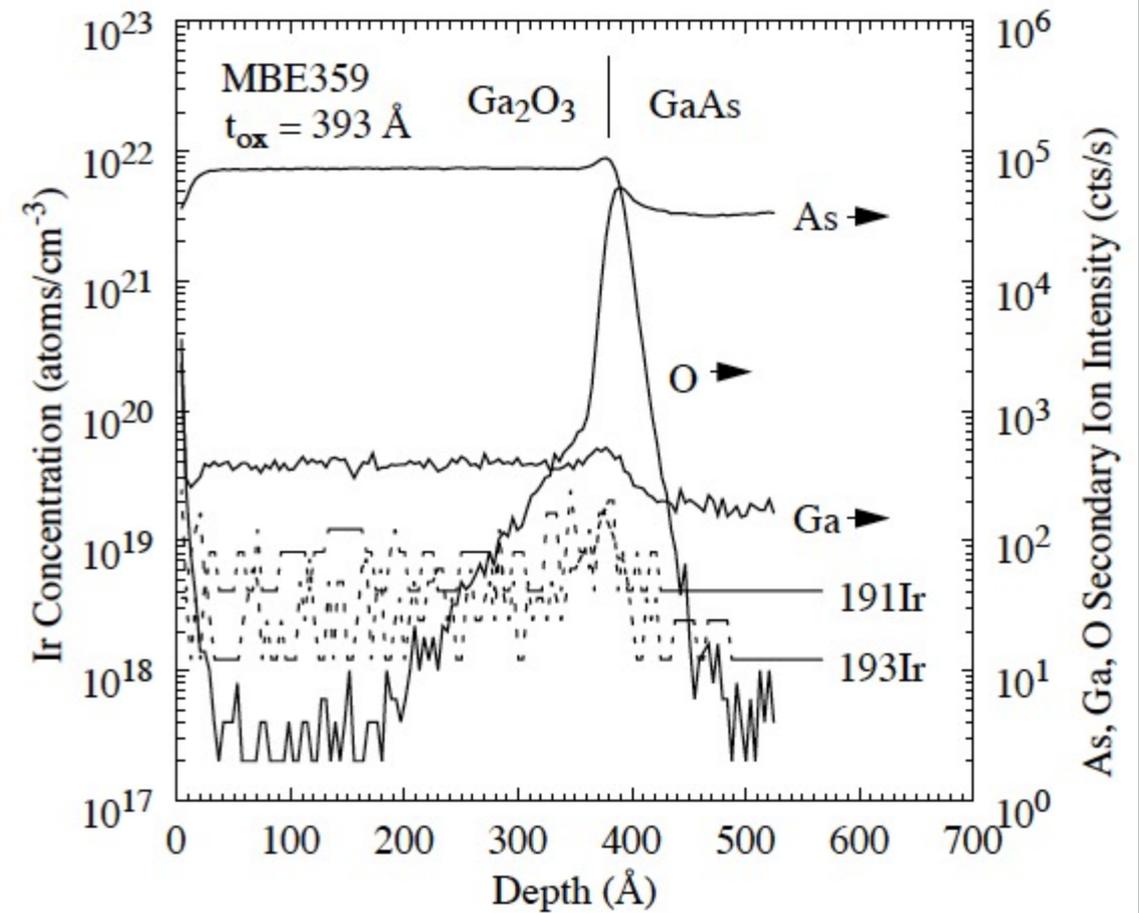
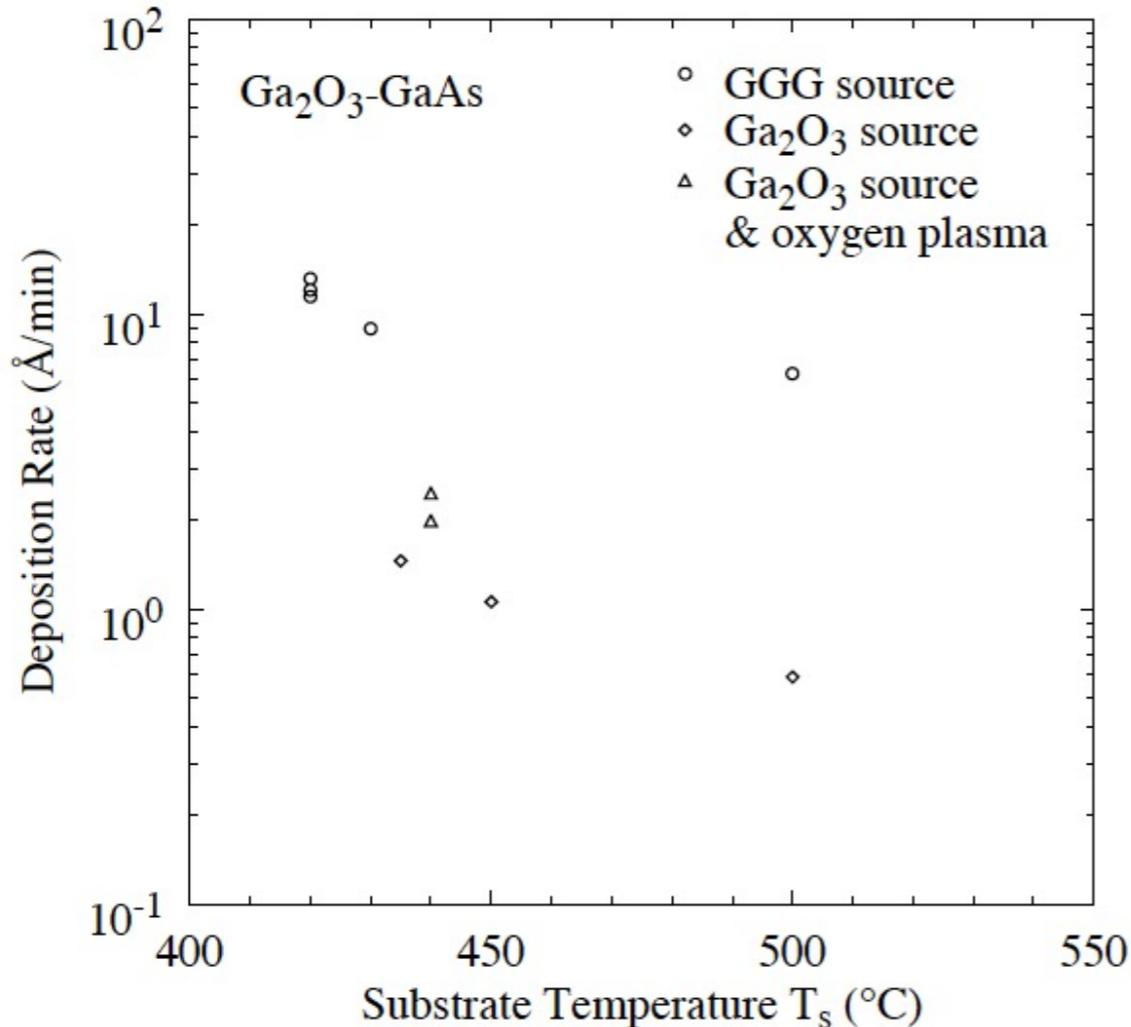


Fig. 9 SIMS depth profile of MBE359. The Ir concentration is in a range of mid 10^{18} cm^{-3} .

Maximum Growth Rate = $0.008 \text{ }\mu\text{m}/\text{hr}$

Ga_2O_3 in Ir crucible at $T = 1800\text{-}1980 \text{ }^\circ\text{C}$

Use Ga(*l*) + Ga₂O₃(s) Mixture?



THE PRESSURE OF Ga₂O OVER GALLIUM-Ga₂O₃ MIXTURES

BY C. J. FROSCH AND C. D. THURMOND

Bell Telephone Laboratories, Inc., Murray Hill, New Jersey

Received October 28, 1961

Mixture with
Ga:Ga₂O₃ = 5:1

TABLE I

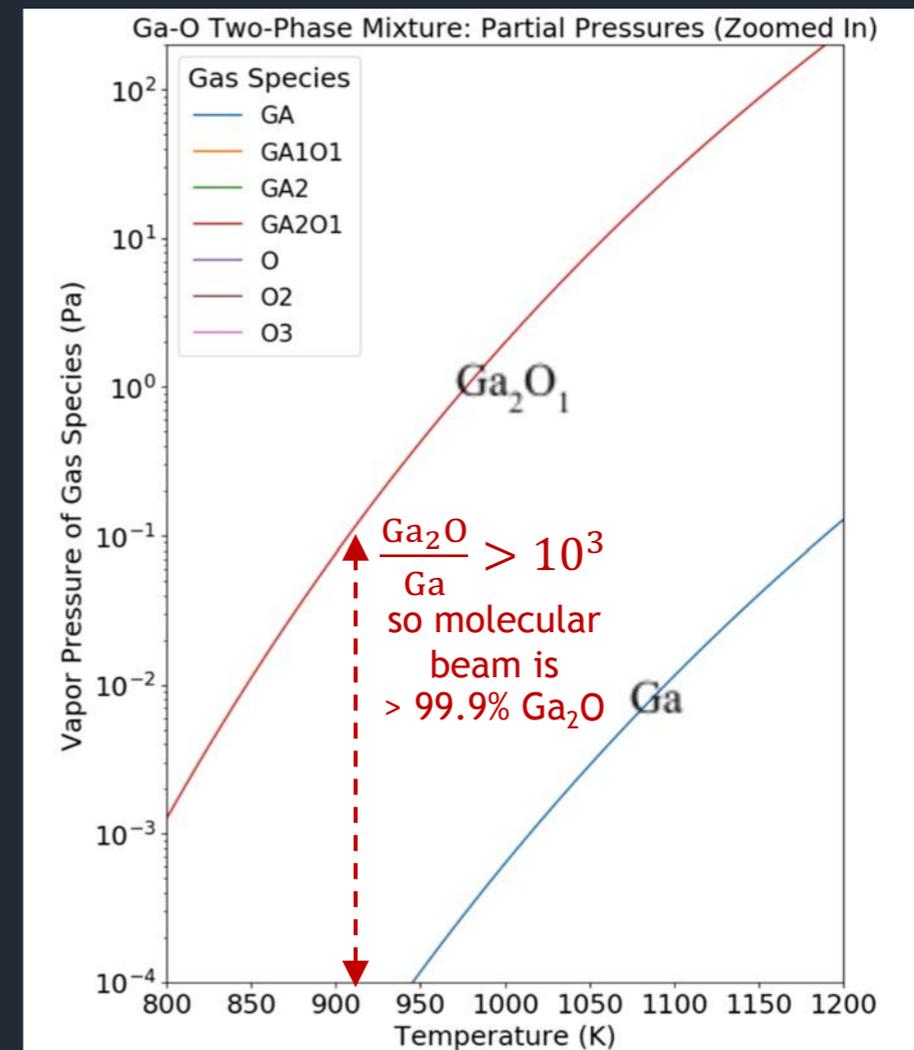
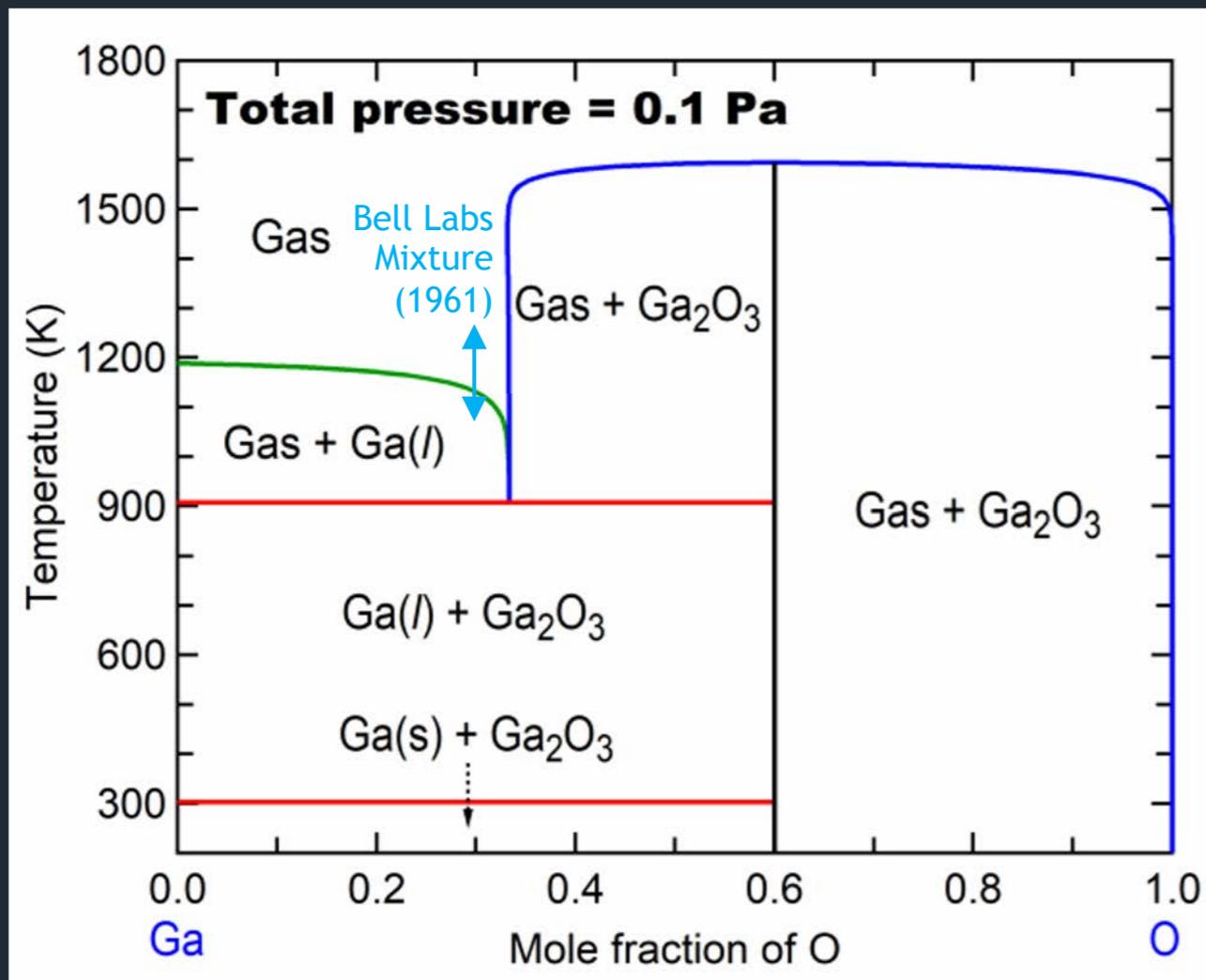
$T^\circ, \text{K.}$	$P_{\text{Ga}_2\text{O}}, \text{atm.}$	ΔH^0_{298}
1073	1.56×10^{-4}	65.7
1173	1.49×10^{-3}	66.3
1223	3.48×10^{-3}	66.7
1273	9.90×10^{-3}	66.6

$$(\Delta H^0_{298}) = 66.3 \text{ kcal./mole}$$

$$(\Delta H^0_{\text{Ga}_2\text{O}})_{298} = -20.7 \text{ kcal./mole}$$

C.J. Frosch and C.D. Thurmond, "The Pressure of Ga₂O over Gallium-Ga₂O₃ Mixtures,"
Journal of Physical Chemistry 66 (1962) 877-878.

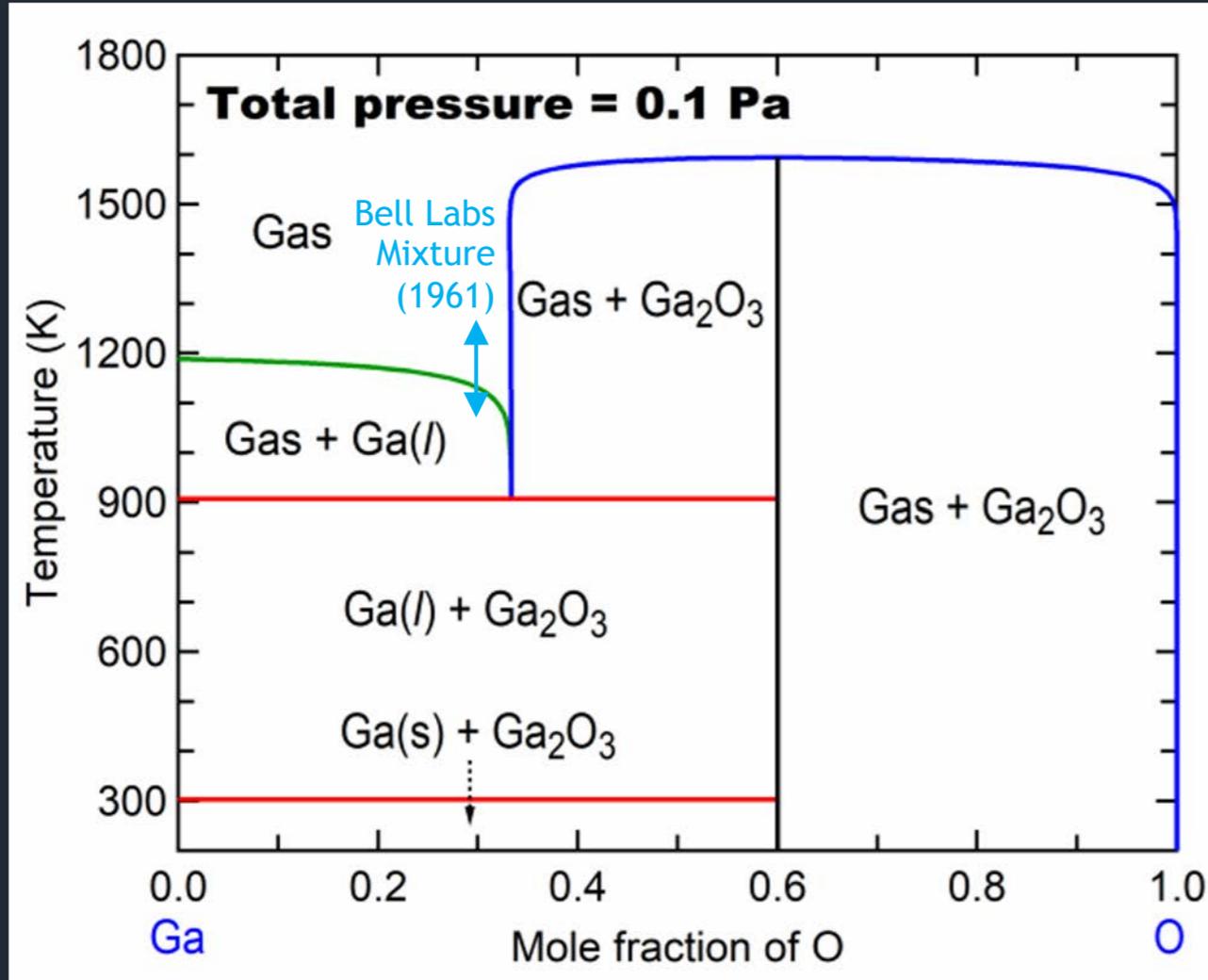
Use Ga(l) + Ga₂O₃(s) Mixture?



P. Vogt, F.V.E. Hensling, K. Azizie, C.S. Chang, D. Turner, J. Park, J.P. McCandless, H. Paik, B.J. Bocklund, G. Hoffman, O. Bierwagen, D. Jena, H.G. Xing, S. Mou, D.A. Muller, S-L. Shang, Z.K. Liu, and D.G. Schlom, "Adsorption-Controlled Growth of Ga₂O₃ by Suboxide Molecular-Beam Epitaxy," *APL Materials* 9 (2021) 031101.

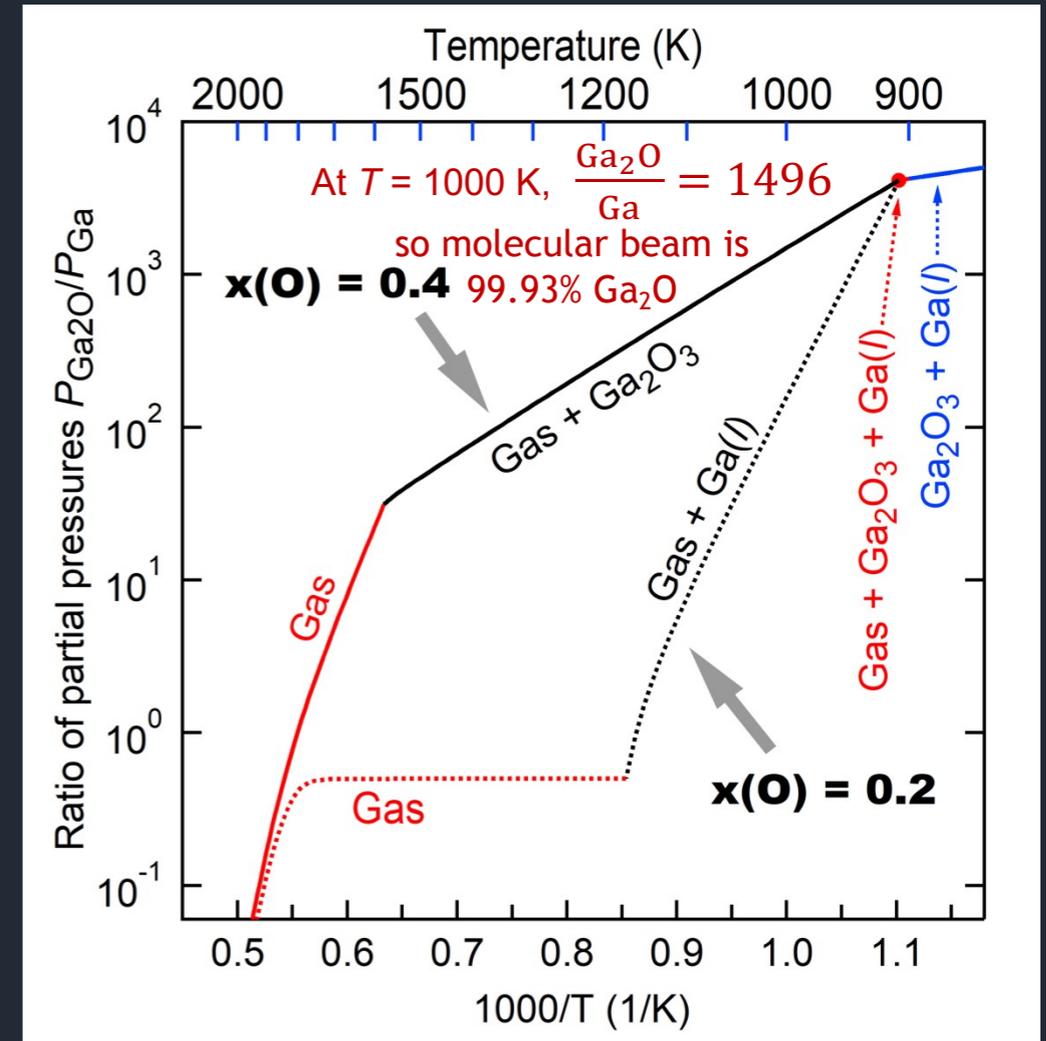
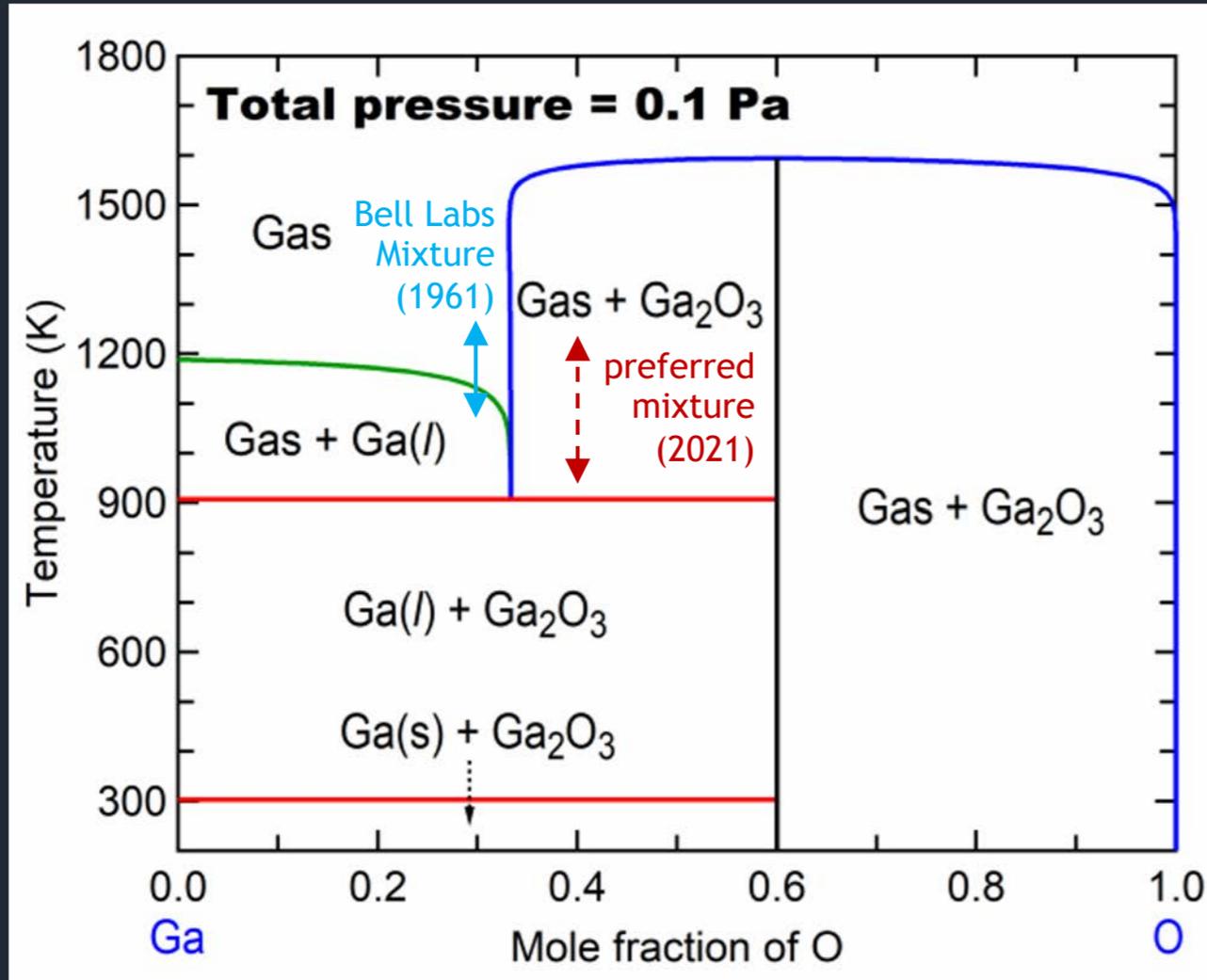
K.M. Adkison, S-L. Shang, B.J. Bocklund, D. Klimm, D.G. Schlom, and Z.K. Liu, "Suitability of Binary Oxides for Molecular-Beam Epitaxy Source Materials: A Comprehensive Thermodynamic Analysis," *APL Materials* 8 (2020) 081110.

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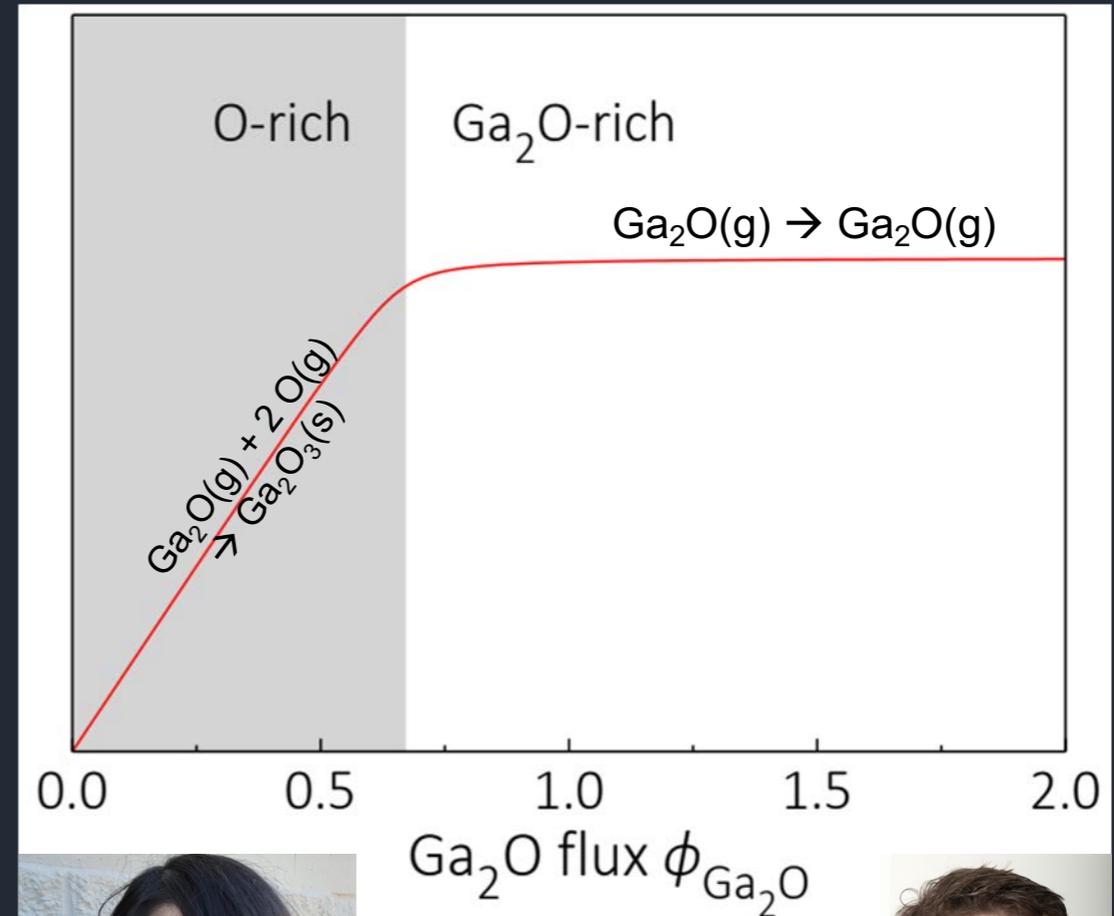
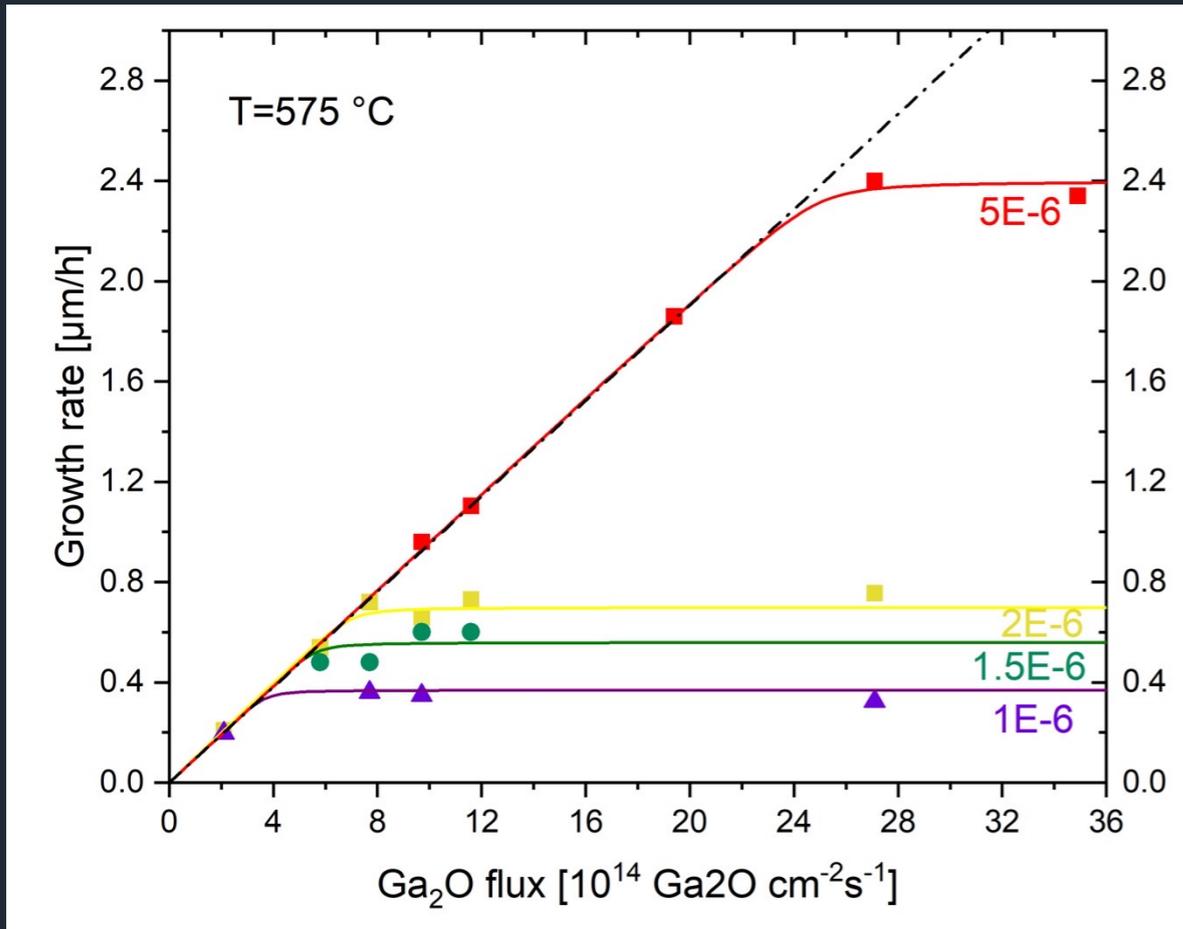
Use Ga(l) + Ga₂O₃(s) Mixture?



Ga₂O₃(s)-rich mixture provides higher Ga₂O/Ga Ratio in molecular beam

P. Vogt, F.V.E. Hensling, K. Azizie, C.S. Chang, D. Turner, J. Park, J.P. McCandless, H. Paik, B.J. Bocklund, G. Hoffman, O. Bierwagen, D. Jena, H.G. Xing, S. Mou, D.A. Muller, S-L. Shang, Z.K. Liu, and D.G. Schlom, "Adsorption-Controlled Growth of Ga₂O₃ by Suboxide Molecular-Beam Epitaxy," *APL Materials* 9 (2021) 031101.

Use Ga(l) + Ga₂O₃(s) Mixture



High growth rate ($> 1 \mu\text{m/hr}$) and epitaxial films at low T_{sub} (e.g., 450°C)

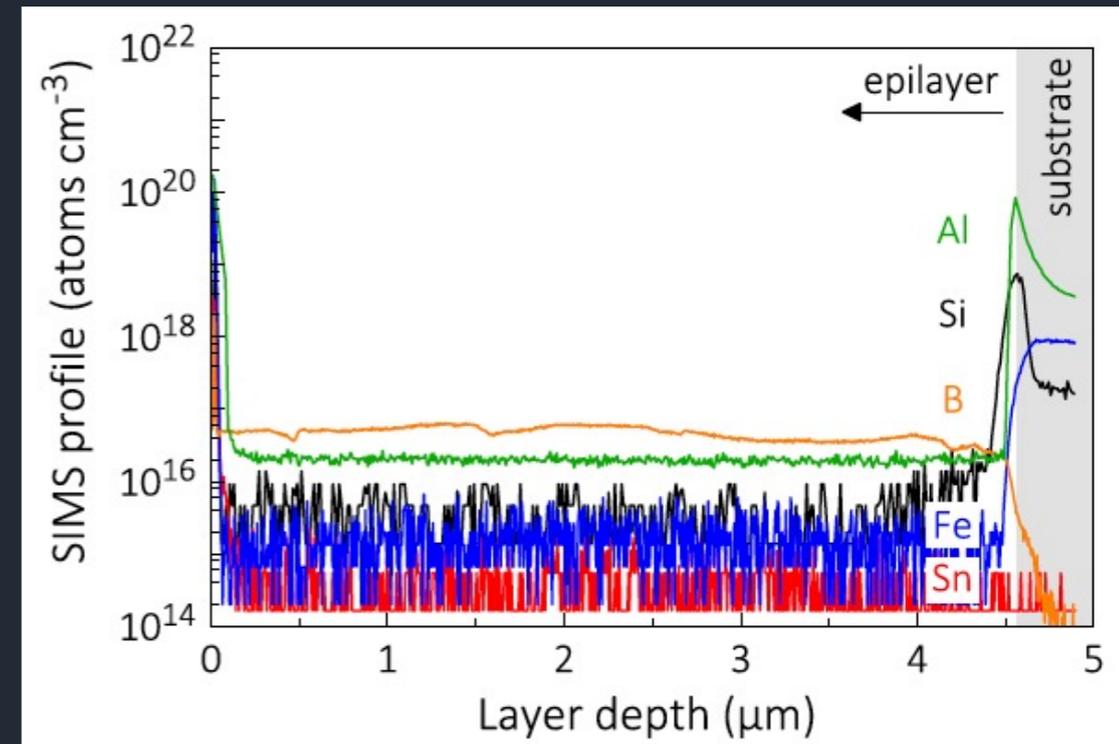
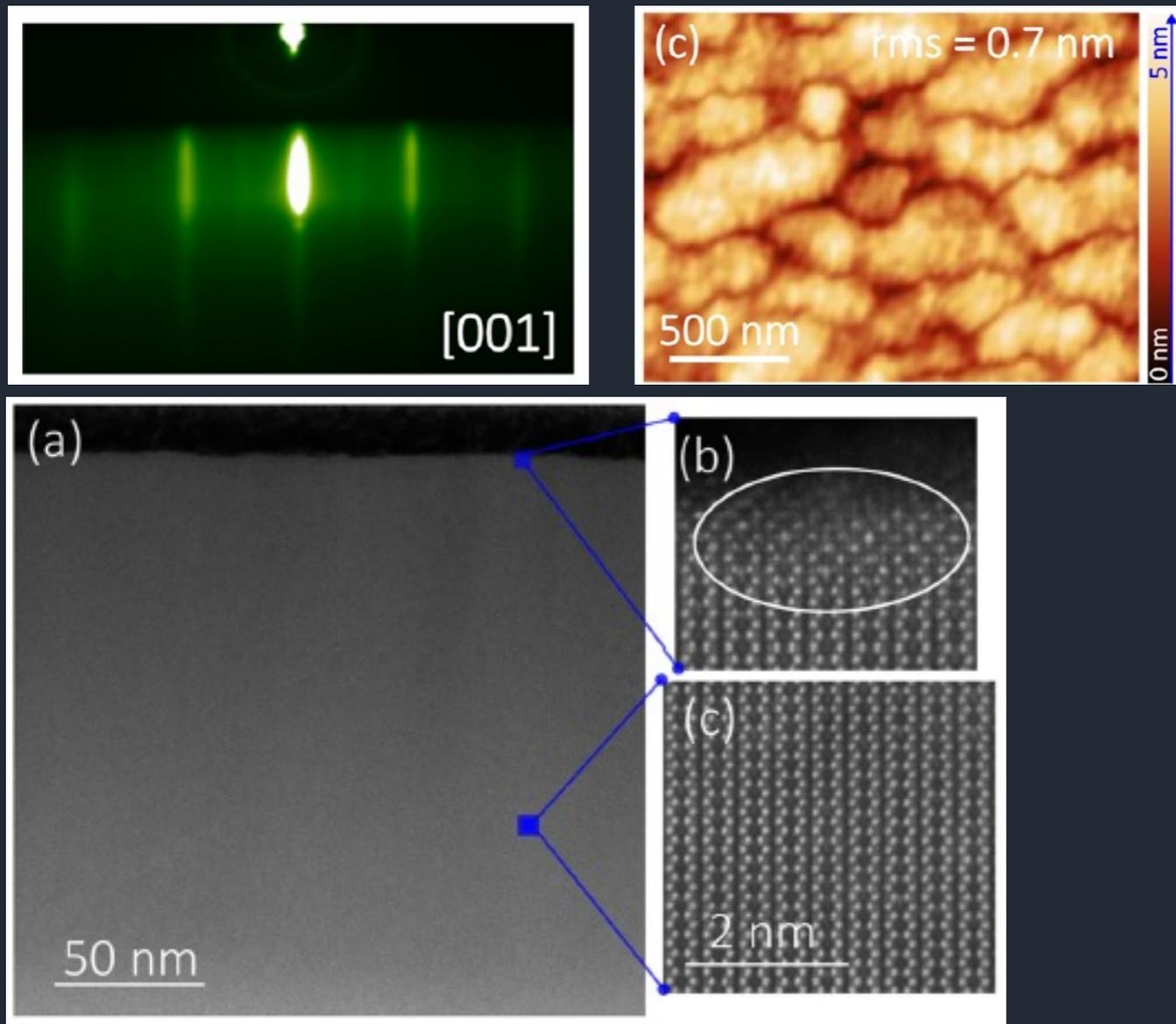


Kathy Azizie



Felix Hensling

Suboxide MBE of β -Ga₂O₃ Grown at 1.2 $\mu\text{m/hr}$



4.5 μm thick film grown at

$$T_{\text{sub}} = 575^\circ \text{C}$$

$$P_{\text{O}_3} = 5 \times 10^{-6} \text{ Torr}$$

Thermo of Suboxide MBE

M_xO_y M_mO_n O M
 ↑? ↑? ↑? ↑?
 Source: M_xO_y

H																				He
Li	Be											B	C	N	O	F				Ne
Na	Mg											Al	Si	P	S	Cl				Ar
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br				Kr
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I				Xe
Cs	Ba		Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At				Rn
Fr	Ra																			
			La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb				Lu
			Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No				Lr

Congruent evaporation from stable source	Incongruent evaporation from metastable source
Incongruent evaporation from stable source	Mainly metal evaporated from metastable source
Mainly metal evaporated from stable source	No thermodynamic data for oxides in SSUB5
Congruent evaporation from metastable source	Not calculated

K.M. Adkison, S-L. Shang, B.J. Bocklund, D. Klimm, D.G. Schlom, and Z.K. Liu, "Suitability of Binary Oxides for Molecular-Beam Epitaxy Source Materials: A Comprehensive Thermodynamic Analysis," *APL Materials* 8 (2020) 081110.

Suboxides offer an alternate means to navigate kinetic pathways

- Comprehensive investigation of vapor pressures of all binary oxides (128 oxides + 27 mixtures)
- 16 evaporate nearly congruently (As_2O_3 , B_2O_3 , BaO , MoO_3 , OsO_4 , P_2O_5 , PbO , PuO_2 , Rb_2O , Re_2O_7 , Sb_2O_3 , SeO_2 , SnO , ThO_2 , Tl_2O , and WO_3)
- + 24 more that could be useful (CeO , Cs_2O , DyO , ErO , Ga_2O , GdO , GeO , HfO , HoO , In_2O , LaO , LuO , NdO , PmO , PrO , PuO , ScO , SiO , SmO , TbO , Te_2O_2 , U_2O_6 , VO_2 , and YO_2)

Materials discovery is more than calculating the properties that a material should have if the atoms were in desired positions. It is also key to get the atoms into those desired positions, to see what the properties really are, and thus realize the potential benefit of a new material. Making this happen takes a combination of ideas, capabilities, and execution—as the recent success by a team led by Assistant and Associate Professors from the University of Michigan illustrates.

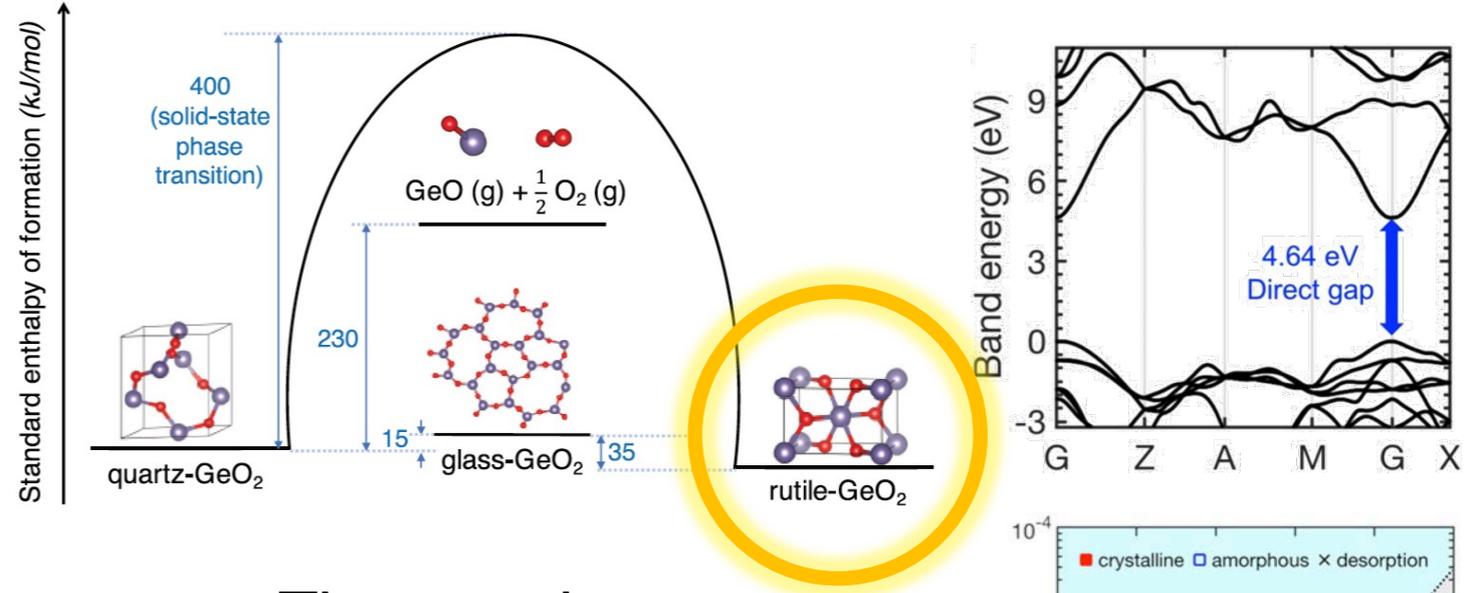
Theoretical work by the team established that rutile-GeO₂—with its ultra-high band gap (4.64 eV), high mobility, high heat conductivity, and desired dopability—could provide superior performance for power electronics. But can this material be made as a thin film? The common synthesis approach would rely on deposition of the constituting elements, but for GeO₂ growth is obstructed by a metastable glass phase and the volatile molecule GeO.

The team came to PARADIM and employed a recently established approach of “sub-oxide MBE”—using partially oxidized GeO instead of Ge—to realize the material in thin film form. Sieun Chae, the same graduate student who did the first-principles calculations, also grew the films. Her work has realized the first single crystal rutile GeO₂ thin films.

S. Chae *et al.* [Appl. Phys. Lett. 117, 072105 \(2020\)](https://doi.org/10.1063/1.5132105).

John T. Heron and Emmanouil Kioupakis, University of Michigan

A promising material is identified by theory



The promise:
high Band Gap ✓
high Mobility ✓
high Thermal Conductivity ✓
But can it be realized?

Adsorption-Controlled Growth of



- **Stannates by Suboxide MBE**

- **BaSnO₃** – H. Paik *et al.*, *APL Materials* 5 (2017) 116107.
- **SnO** – A.B. Mei *et al.*, *Phys. Rev. Mater.* 3 (2019) 105202.
- **Sr₃SnO** – Y. Ma *et al.* *Adv. Mater.* 32 (2020) 2000809.
- **Ta₂SnO₆** – M. Barone *et al.* *J. Phys. Chem. C* 126 (2022) 3764–3775.

- **Gallates by Suboxide MBE**

- **Ga₂O₃** – P. Vogt *et al.*, *APL Mater.* 9 (2021) 031101.

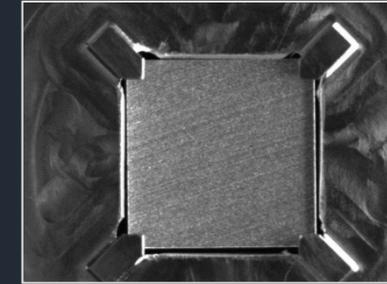
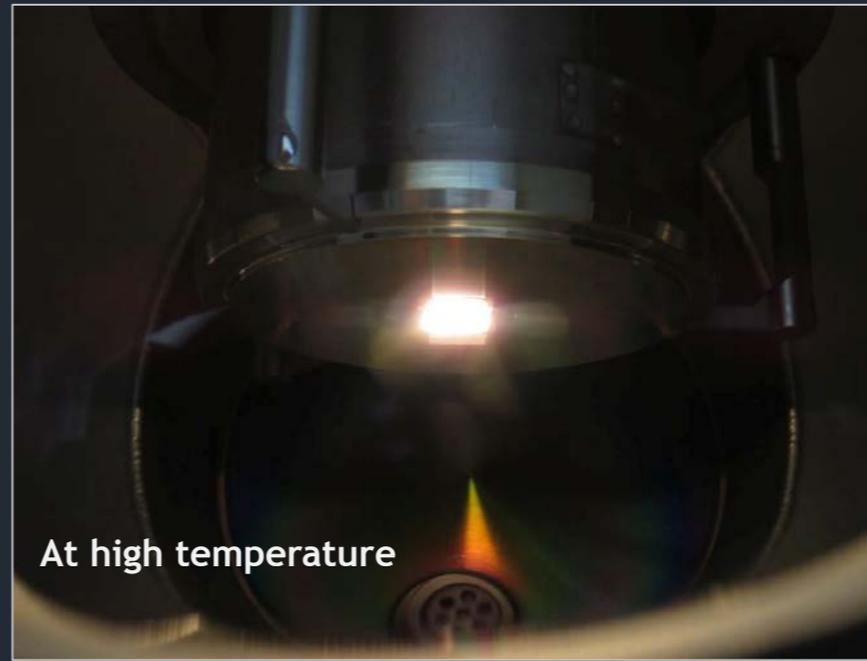
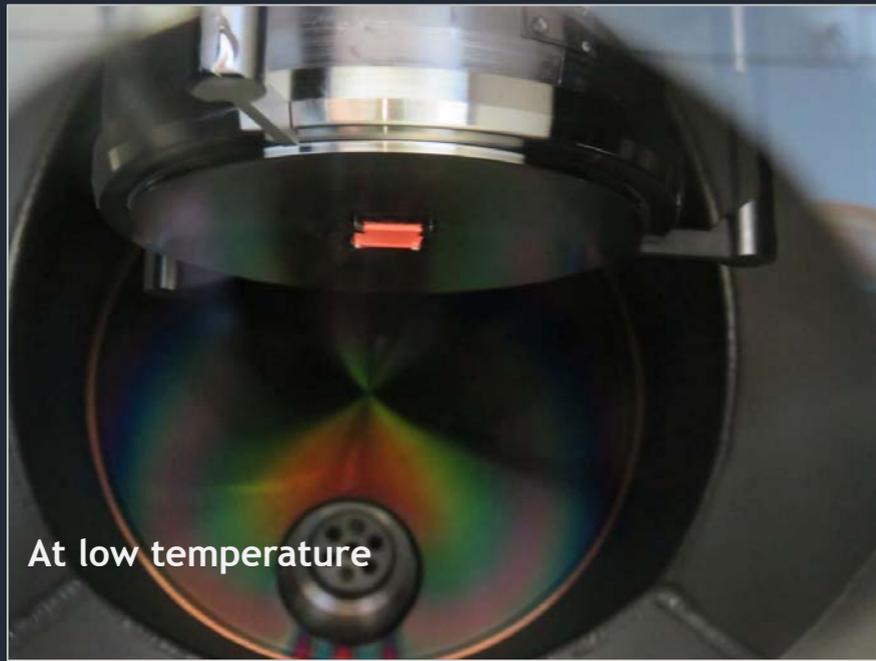
- **Indates by Suboxide MBE**

- **In₂O₃** – P. Vogt *et al.*, *Phys. Rev. Appl.* 17 (2022) 034021 .

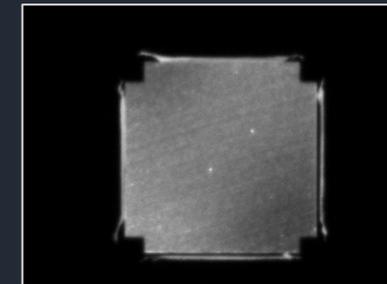
Laser Substrate Heater

CO₂ Laser Substrate Heater

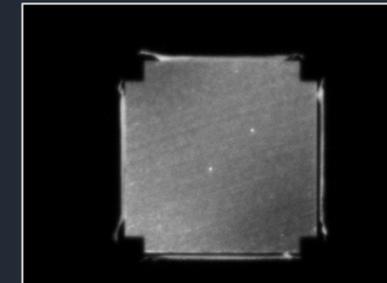
- From EpiRay <https://epiray.de>
- Currently being tested
- Expect On-Line in PARADIM by end of 2022



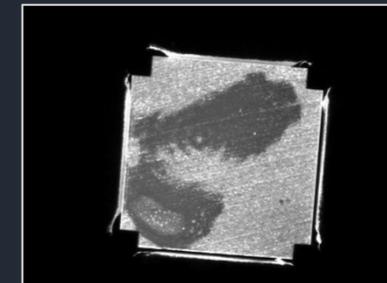
Laser off



Homogeneity
at 1000 ° C



Homogeneity
at 1800 ° C



Laser power 95W
Melting starts
(2040 ° C)

Laser Substrate Heater

- Laser heater for MBE

- T_{sub} up to 2000 °C
- *In situ* substrate termination demonstrated for:

MgO

Al₂O₃

SrTiO₃

LaAlO₃

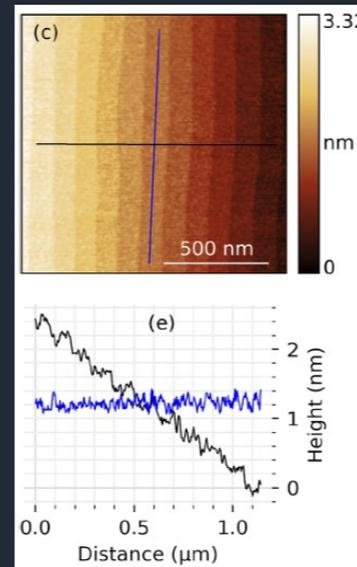
NdGaO₃

DyScO₃

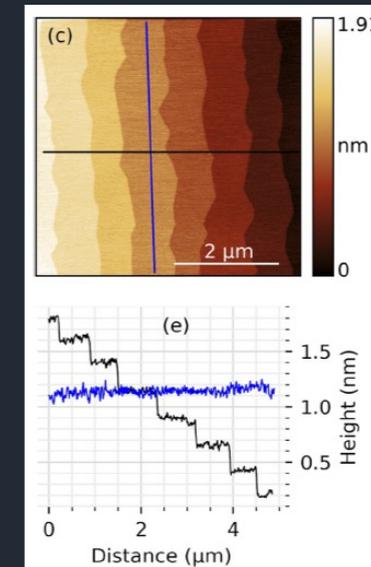
TbScO₃

Expect On-Line in
PARADIM by
end of 2022

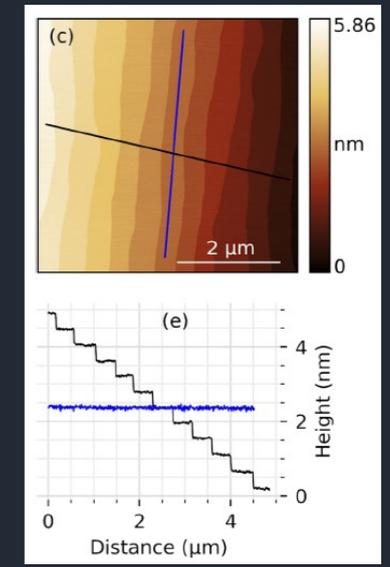
MgO at 1773 °C



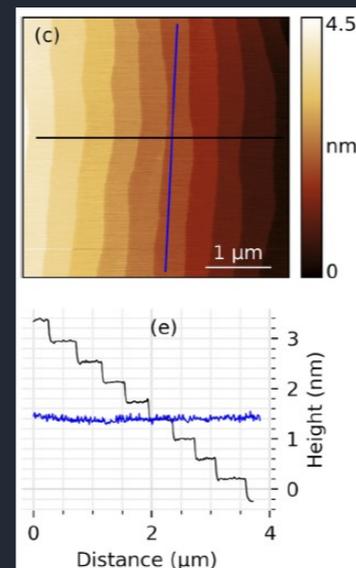
Al₂O₃ at 1533 °C



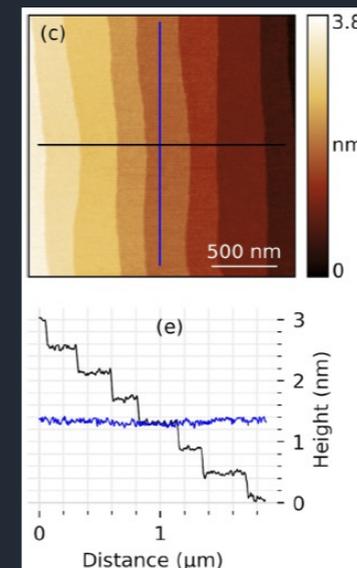
SrTiO₃ at 1460 °C



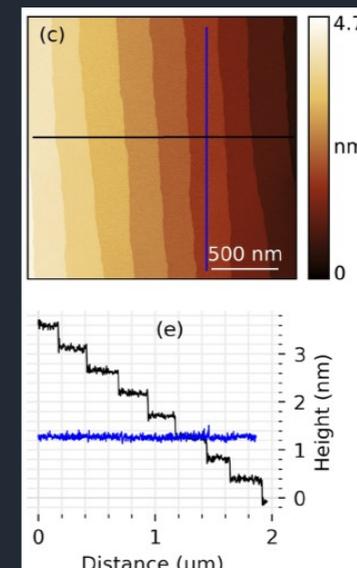
LaAlO₃ at 1565 °C



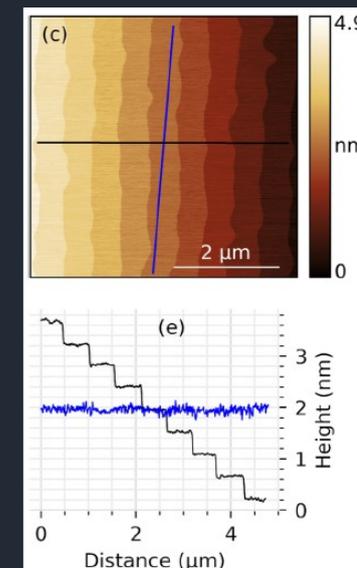
NdGaO₃ at 1043 °C



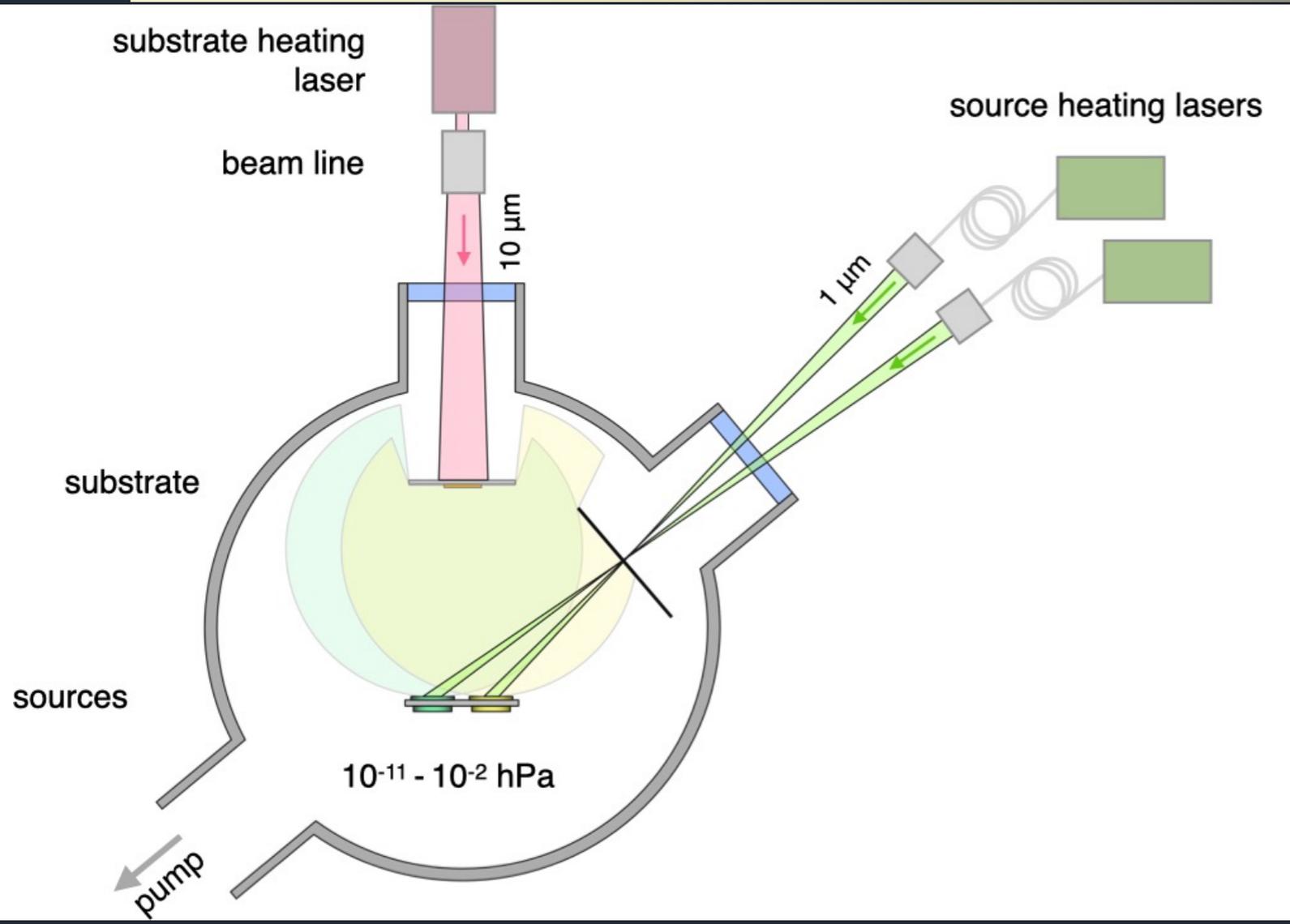
DyScO₃ at 1460 °C



TbScO₃ at 1356 °C



Future of Oxide MBE



Expanded Growth Conditions
Unlimited temperatures for substrate and sources

Robust
All heaters *outside* vacuum

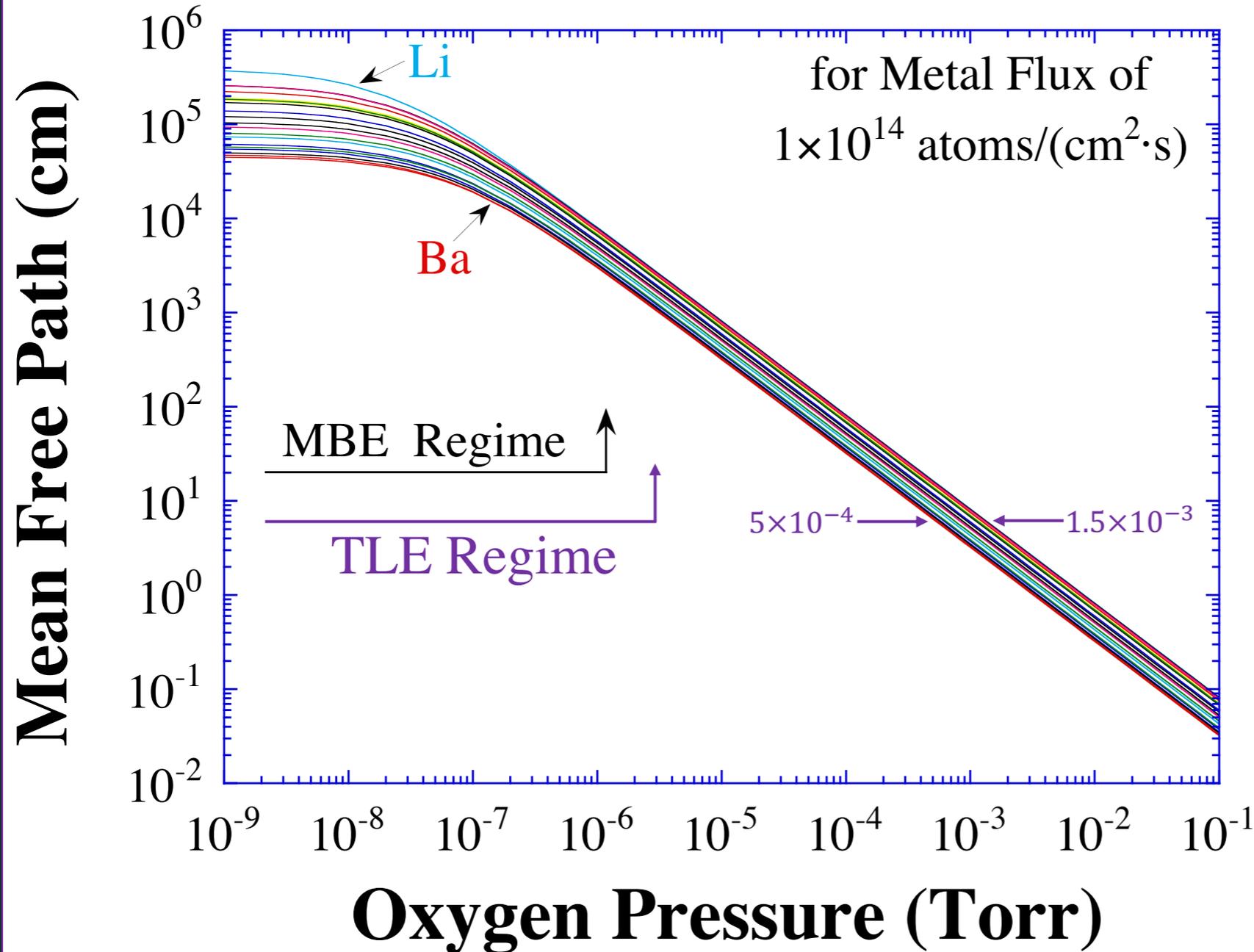
Achieve Oxidation of the Film
O₃ up to 10⁻³ Torr

Prevent Problematic Oxidation of the Sources
Surface continuously evaporated

Substrate Termination
In situ thermal termination

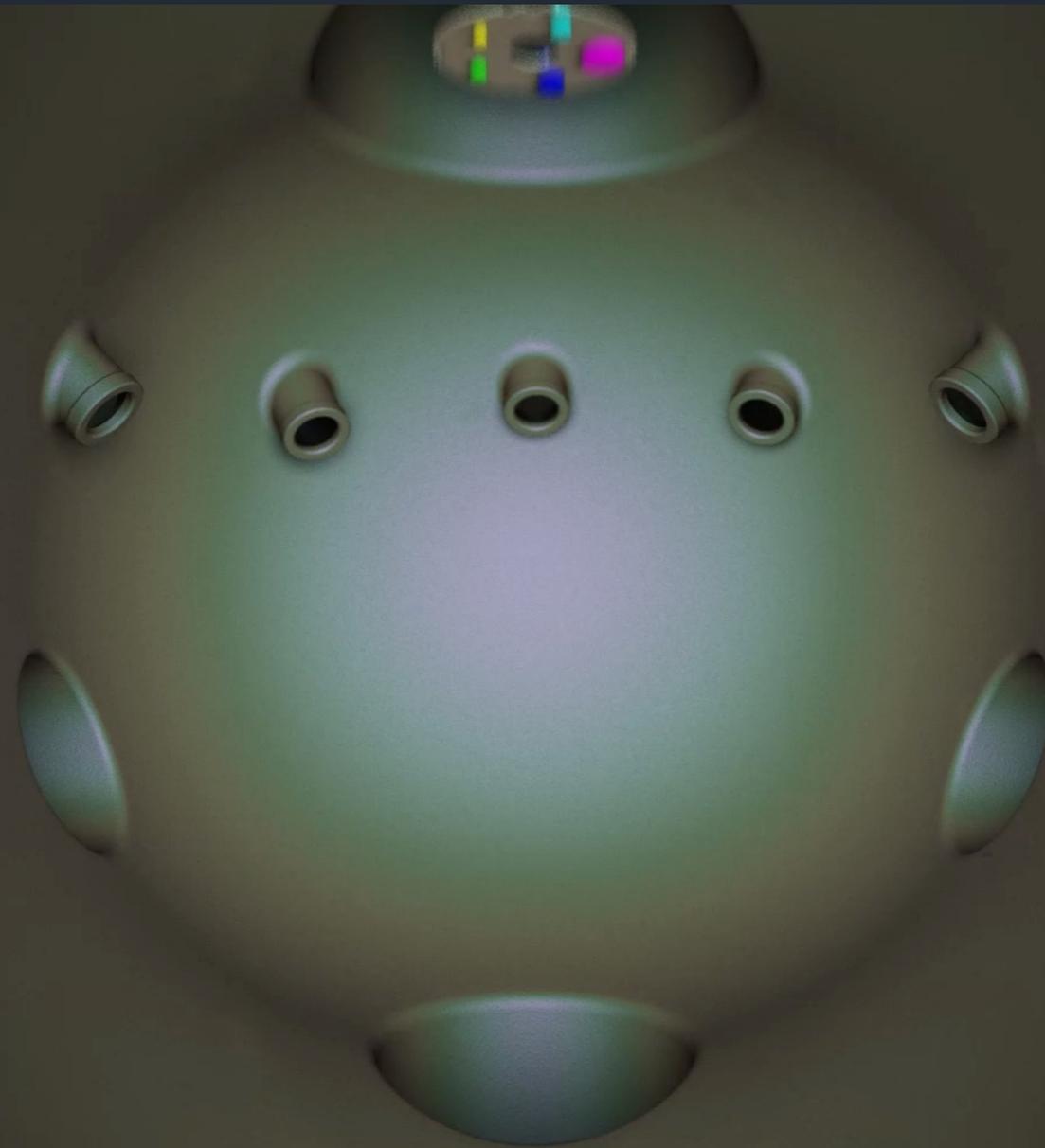
W. Braun and J. Mannhart,
“Film Deposition by **Thermal Laser Evaporation**,”
AIP Advances 9 (2019) 085310.

Maximum O₂ Pressure for MBE TLE?



D.G. Schlom and J.S. Harris, Jr., in *Molecular Beam Epitaxy: Applications to Key Materials*, edited by R.F.C. Farrow (Noyes, Park Ridge, 1995), pp. 505-622.

TLE — since 2019



Video courtesy
of
Epiray GmbH

GOT MBE?



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